

Determination of Molar Mass by Vapor Density

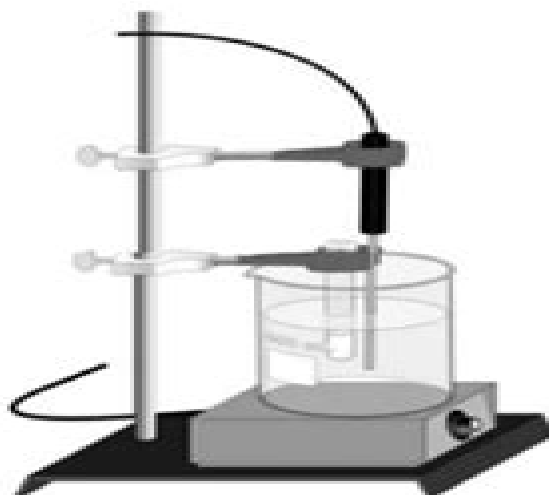
One of the properties that helps characterize a substance is its molar mass. If the substance in question is a volatile liquid, a common method to determine its molar mass is to vaporize it and apply the ideal gas law, $PV = nRT$ to the data collected. Because the liquid is *volatile*, it can easily be converted to a vapor. Volatile substances are usually composed of nonpolar molecules. As a result the molecules have primarily London dispersion forces and very little thermal energy is required to overcome these attractive forces since the molecules are relatively small. Therefore, the liquid vaporizes easily and quickly at temperatures less than 100°C. While the substance is in the vapor phase, you can measure its volume, pressure, and temperature. You can then use the ideal gas law to calculate the number of moles of the substance. Finally, you can use the number of moles of the gas to calculate molar mass.

OBJECTIVES

In this experiment, you will

- Evaporate a sample of a liquid substance and measure certain physical properties of the substance as it condenses.
- Determine the molar mass of an unknown liquid.

Best results are obtained when the test tube containing the sample is submerged in the water bath to just below the foil cap. Consider using tall form beakers for your water baths.



Determination Of Molar Mass By Vapor Density

David Clark



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