

SAMPLE EXERCISE 3.15 Determining Empirical Formula by Combustion Analysis

Isopropyl alcohol, a substance sold as rubbing alcohol, is composed of C, H, and O. Combustion of 0.255 g of isopropyl alcohol produces 0.561 g of CO_2 and 0.306 g of H_2O . Determine the empirical formula of isopropyl alcohol.

Solution

Analyze: We are told that isopropyl alcohol contains C, H, and O atoms and given the quantities of CO_2 and H_2O produced when a given quantity of the alcohol is combusted. We must use this information to determine the empirical formula for isopropyl alcohol, a task that requires us to calculate the number of moles of C, H, and O in the sample.

Plan: We can use the mole concept to calculate the number of grams of C present in the CO_2 and the number of grams of H present in the H_2O . These are the quantities of C and H present in the isopropyl alcohol before combustion. The number of grams of O in the compound equals the mass of the isopropyl alcohol minus the sum of the C and H masses. Once we have the number of grams of C, H, and O in the sample, we can then proceed as in Sample Exercise 3.13: Calculate the number of moles of each element, and determine the mole ratio, which gives the subscripts in the empirical formula.

Solve: To calculate the number of grams of C, we first use the molar mass of CO_2 , $1 \text{ mol CO}_2 = 44.0 \text{ g CO}_2$, to convert grams of CO_2 to moles of CO_2 . Because there is only 1 C atom in each CO_2 molecule, there is 1 mol of C atoms per mole of CO_2 molecules. This fact allows us to convert the moles of CO_2 to moles of C. Finally, we use the molar mass of C, $1 \text{ mol C} = 12.0 \text{ g C}$, to convert moles of C to grams of C. Combining the three conversion factors, we have:

$$\text{Grams C} = (0.561 \text{ g CO}_2) \left(\frac{1 \text{ mol CO}_2}{44.0 \text{ g CO}_2} \right) \left(\frac{1 \text{ mol C}}{1 \text{ mol CO}_2} \right) \left(\frac{12.0 \text{ g C}}{1 \text{ mol C}} \right) = 0.153 \text{ g C}$$

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CO Houle



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