

## Determining Empirical Formulas

1) 75% carbon, 25% hydrogen

Answer:

 $\text{CH}_4$ 

Work:

$$\text{C } 75\% \rightarrow 75\text{g C} \left| \frac{1 \text{ mol}}{12.01\text{g C}} \right| = 6.24 \text{ mol} / 6.24 = 1$$

$$\text{H } 25\% \rightarrow 25\text{g H} \left| \frac{1 \text{ mol H}}{1.01\text{g H}} \right| = 24.75 \text{ mol} / 6.24 = 4$$



2) 52.7% potassium, 47.3% chlorine

Answer:

 $\text{KCl}$ 

Work:

$$\text{K } 52.7\% \rightarrow 52.7\text{g K} \left| \frac{1 \text{ mol K}}{39.10\text{g K}} \right| = 1.34 \text{ mol} / 1.33 = 1$$

$$\text{Cl } 47.3\% \rightarrow 47.3\text{g Cl} \left| \frac{1 \text{ mol Cl}}{35.45\text{g Cl}} \right| = 1.33 \text{ mol} / 1.33 = 1$$



3) 22.1% aluminum, 25.4% phosphorous, 52.5 % oxygen

Answer:

 $\text{AlPO}_4$ 

Work:

$$\text{Al } 22.1\% \rightarrow 22.1\text{g Al} \left| \frac{1 \text{ mol Al}}{26.98\text{g Al}} \right| = 0.819 \text{ mol} / 0.819 = 1$$

$$\text{P } 25.4\% \rightarrow 25.4\text{g P} \left| \frac{1 \text{ mol P}}{30.97\text{g P}} \right| = 0.820 \text{ mol} / 0.819 = 1$$

$$\text{O } 52.5\% \rightarrow 52.5\text{g O} \left| \frac{1 \text{ mol O}}{16.00\text{g O}} \right| = 3.28 \text{ mol} / 0.819 = 4$$



4) 13% magnesium, 87% bromine

Answer:

 $\text{MgBr}_2$ 

Work:

$$13\text{g Mg} \left| \frac{1 \text{ mol Mg}}{24.31\text{g Mg}} \right| = 0.54 \text{ mol Mg} / 0.54 = 1$$

$$87\text{g Br} \left| \frac{1 \text{ mol Br}}{79.90\text{g Br}} \right| = 1.09 \text{ mol Br} / 0.54 = 2$$



5) 32.4% sodium, 22.5% sulfur, 45.1% oxygen

Answer:

 $\text{Na}_2\text{SO}_4$ 

Work:

$$32.4\text{g Na} \left| \frac{1 \text{ mol}}{22.99\text{g Na}} \right| = 1.41 \text{ mol Na} / 0.70 = 2$$

$$22.5\text{g S} \left| \frac{1 \text{ mol S}}{32.37\text{g S}} \right| = 0.70 \text{ mol S} / 0.70 = 1$$

$$45.1\text{g O} \left| \frac{1 \text{ mol O}}{16\text{g O}} \right| = 2.82 \text{ mol O} / 0.70 = 4$$



# Determining Chemical Formulas Answer Key

**Brendan G. Carr**



## **Determining Chemical Formulas Answer Key:**

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An Introduction to Chemistry Michael Mosher,Paul Kelter,2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

**eBook: General, Organic and Biological Chemistry 2e** SMITH,2012-02-16 eBook

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