

Example 1: Calculate the empirical formula for a substance containing 1.11 g of Se and 1.61 g of F.

$$1.11 \text{ g Se} \times \frac{1 \text{ mol Se}}{78.971 \text{ g Se}} = 0.0141 \text{ mol Se}$$

$$1.61 \text{ g F} \times \frac{1 \text{ mol F}}{18.998 \text{ g F}} = 0.0847 \text{ mol F}$$

$$\text{Se} \frac{0.0141}{0.0141} \text{ F} \frac{0.0847}{0.0141} = \text{Se F}_6$$

Determining Empirical Formulas 55

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