

Determining Empirical Formulas

1) 75% carbon, 25% hydrogen	Answer:	CH ₄		
Work:				
C 75% → 75g C $\frac{1 \text{ mol}}{12.01 \text{ g C}} = 6.24 \text{ mol} / 6.24 = 1$	<div style="border: 1px solid black; padding: 5px; display: inline-block;">CH₄</div>			
H 25% → 25g H $\frac{1 \text{ mol H}}{1.01 \text{ g H}} = 24.75 \text{ mol} / 6.24 = 4$				
2) 52.7% potassium, 47.3% chlorine	Answer:	KCl		
Work:				
K 52.7% → 52.7g K $\frac{1 \text{ mol K}}{39.10 \text{ g K}} = 1.34 \text{ mol} / 1.33 = 1$	<div style="border: 1px solid black; padding: 5px; display: inline-block;">KCl</div>			
Cl 47.3% → 47.3g Cl $\frac{1 \text{ mol Cl}}{35.45 \text{ g Cl}} = 1.33 \text{ mol} / 1.33 = 1$				
3) 22.1% aluminum, 25.4% phosphorous, 52.5 % oxygen	Answer:	AlPO ₄		
Work:				
Al 22.1% → 22.1g Al $\frac{1 \text{ mol Al}}{26.98 \text{ g Al}} = 0.819 \text{ mol} / 0.819 = 1$	<div style="border: 1px solid black; padding: 5px; display: inline-block;">AlPO₄</div>			
P 25.4% → 25.4g P $\frac{1 \text{ mol P}}{30.97 \text{ g P}} = 0.820 \text{ mol} / 0.819 = 1$				
O 52.5% → 52.5g O $\frac{1 \text{ mol O}}{16.00 \text{ g O}} = 3.28 \text{ mol} / 0.819 = 4$				
4) 13% magnesium, 87% bromine	Answer:	MgBr ₂		
Work:				
13g Mg $\frac{1 \text{ mol Mg}}{24.31 \text{ g Mg}} = 0.54 \text{ mol Mg} / 0.54 = 1$	<div style="border: 1px solid black; padding: 5px; display: inline-block;">MgBr₂</div>			
87g Br $\frac{1 \text{ mol Br}}{79.90 \text{ g Br}} = 1.09 \text{ mol Br} / 0.54 = 2$				
5) 32.4% sodium, 22.5% sulfur, 45.1% oxygen	Answer:	Na ₂ SO ₄		
Work:				
32.4g Na $\frac{1 \text{ mol Na}}{22.99 \text{ g Na}} = 1.41 \text{ mol Na} / 0.70 = 2$	<div style="border: 1px solid black; padding: 5px; display: inline-block;">Na₂SO₄</div>			
22.5g S $\frac{1 \text{ mol S}}{32.07 \text{ g S}} = 0.70 \text{ mol S} / 0.70 = 1$				
45.1g O $\frac{1 \text{ mol O}}{16.00 \text{ g O}} = 2.82 \text{ mol O} / 0.70 = 4$				

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Ian Pickup



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Table of Contents Determining Empirical Formulas Worksheet Chemistry If8766 Answers

1. Understanding the eBook Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - The Rise of Digital Reading Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - Personalized Recommendations
 - Determining Empirical Formulas Worksheet Chemistry If8766 Answers User Reviews and Ratings
 - Determining Empirical Formulas Worksheet Chemistry If8766 Answers and Bestseller Lists
5. Accessing Determining Empirical Formulas Worksheet Chemistry If8766 Answers Free and Paid eBooks
 - Determining Empirical Formulas Worksheet Chemistry If8766 Answers Public Domain eBooks
 - Determining Empirical Formulas Worksheet Chemistry If8766 Answers eBook Subscription Services
 - Determining Empirical Formulas Worksheet Chemistry If8766 Answers Budget-Friendly Options
6. Navigating Determining Empirical Formulas Worksheet Chemistry If8766 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Determining Empirical Formulas Worksheet Chemistry If8766 Answers Compatibility with Devices
 - Determining Empirical Formulas Worksheet Chemistry If8766 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - Highlighting and Note-Taking Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - Interactive Elements Determining Empirical Formulas Worksheet Chemistry If8766 Answers
8. Staying Engaged with Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Determining Empirical Formulas Worksheet Chemistry If8766 Answers
9. Balancing eBooks and Physical Books Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Determining Empirical Formulas Worksheet Chemistry If8766 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - Setting Reading Goals Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - Fact-Checking eBook Content of Determining Empirical Formulas Worksheet Chemistry If8766 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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