

1. An unknown compound was found to have a percent composition as follows: 47.0% potassium, 14.5% carbon, and 38.5% oxygen. What is its empirical formula?

$$\begin{aligned}
 \text{K} &= 47\% \left(\frac{1}{39}\right) = 1.205 \cancel{12} = 1 \\
 \text{C} &= 14.5\% \left(\frac{1}{12}\right) = 1.208 \cancel{12} = 1 \\
 \text{O} &= 38.5\% \left(\frac{1}{16}\right) = 2.406 \cancel{12} = 2
 \end{aligned}
 \quad \boxed{\text{KCO}_2}$$

2. A sample of iodine chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the iodine compound?

$$\begin{aligned}
 \text{I} &= 0.2596 \left(\frac{1}{127}\right) = 0.002044 \cancel{0.0022} = 1 \\
 \text{Cl} &= 0.2404 \left(\frac{1}{35.5}\right) = 0.00677 \cancel{0.0022} = 3
 \end{aligned}
 \quad \boxed{\text{I}_2\text{Cl}_3}$$

3. Nitrogen and oxygen form an extensive series of oxides with the general formula N_xO_y . One of them is a blue solid that comes apart, reversibly, in the gas phase. It contains 36.84% N. What is the empirical formula of this oxide?

$$\begin{aligned}
 \text{N} &= 36.84 \left(\frac{1}{14}\right) = 2.631 \cancel{2.65} = 1 + 2 = 2 \\
 \text{O} &= 63.16 \left(\frac{1}{16}\right) = 3.9475 \cancel{2.65} = 1.5 + 2 = 3
 \end{aligned}
 \quad \boxed{\text{N}_2\text{O}_3}$$

4. Rubbing alcohol was found to contain 60.0% carbon, 13.4% hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of rubbing alcohol?

$$\begin{aligned}
 \text{C} &= 60 \left(\frac{1}{12}\right) = 5 \text{ Atoms} = 5 \\
 \text{H} &= 13.4 \left(\frac{1}{1}\right) = 13.4 \cancel{13.4} = 7 \\
 \text{O} &= 26.6 \left(\frac{1}{16}\right) = 1.66 \cancel{1.66} = 1
 \end{aligned}
 \quad \boxed{\text{C}_5\text{H}_7\text{O}}$$

5. What is the empirical formula of a compound that contains 0.783g of Carbon, 0.186g of Hydrogen and 0.521g of Oxygen?

$$\begin{aligned}
 \text{C} &= 0.783 \left(\frac{1}{12}\right) = 0.06525 \cancel{0.065} = 2 \\
 \text{H} &= 0.186 \left(\frac{1}{1}\right) = 0.186 \cancel{0.065} = 6 \\
 \text{O} &= 0.521 \left(\frac{1}{16}\right) = 0.03256 \cancel{0.0325} = 1
 \end{aligned}
 \quad \boxed{\text{C}_2\text{H}_6\text{O}}$$

6. What is empirical formula of a compound which consists of 89.14% Au and 10.86% of O?

$$\begin{aligned}
 \text{Au} &= 89.14 \left(\frac{1}{197}\right) = 0.452 \cancel{0.42} = 1 + 2 = 2 \\
 \text{O} &= 10.8 \left(\frac{1}{16}\right) = 0.675 \cancel{0.42} = 1.5 + 2 = 3
 \end{aligned}
 \quad \boxed{\text{Au}_2\text{O}_3}$$

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RS Peters



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