

DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Name Max Lundy

Solve the problems below.

1. The empirical formula of a compound is NO_2 . Its molecular mass is 92 g/mol. What is its molecular formula?

$$\begin{aligned} (1) \quad & \frac{14 + 2(16)}{46} = 2 \\ & 92 \div 46 = 2 \\ & 2(\text{N}) 2(\text{O}_2) \\ & \text{N}_2\text{O}_4 \end{aligned}$$



2. The empirical formula of a compound is CH_2 . Its molecular mass is 70 g/mol. What is its molecular formula?

$$\begin{aligned} (1) \quad & \frac{12 + 2}{14} = 5 \\ & 70 \div 14 = 5 \\ & 5(\text{C}) 5(\text{H}_2) \\ & \text{C}_5\text{H}_{10} \end{aligned}$$



3. A compound is found to be 40.0% carbon, 6.7% hydrogen and 53.3% oxygen. Its molecular mass is 60 g/mol. What is its molecular formula?

$$\begin{aligned} (1) \quad & 40 \left(\frac{1}{12} \right) = 3.333 \quad \frac{60}{12} = 5 \quad \text{C}_5\text{H}_{10}\text{O} \\ (2) \quad & 6.7 \left(\frac{1}{1} \right) = 6.7 \quad \frac{60}{1} = 60 \quad 2(\text{C}) 2(\text{H}_6) 2(\text{O}) \\ (3) \quad & 53.3 \left(\frac{1}{16} \right) = 3.333 \quad \frac{60}{16} = 3.75 \quad \text{C}_2\text{H}_4\text{O}_2 \end{aligned}$$



4. A compound is 64.9% carbon, 13.5% hydrogen and 21.6% oxygen. Its molecular mass is 74 g/mol. What is its molecular formula?

$$\begin{aligned} (1) \quad & 64.9 \left(\frac{1}{12} \right) = 5.4083 \quad \frac{74}{12} = 6.166 \quad \text{C}_4\text{H}_{10}\text{O} \\ (2) \quad & 13.5 \left(\frac{1}{1} \right) = 13.5 \quad \frac{74}{1} = 74 \quad 4(12) + 10 + 16 = 74 \\ (3) \quad & 21.6 \left(\frac{1}{16} \right) = 1.35 \quad \frac{74}{16} = 4.625 \quad \text{C}_4\text{H}_{10}\text{O} \end{aligned}$$



5. A compound is 54.5% carbon, 9.1% hydrogen and 36.4% oxygen. Its molecular mass is 88 g/mol. What is its molecular formula?

$$\begin{aligned} (1) \quad & 54.5 \left(\frac{1}{12} \right) = 4.5416 \quad \frac{88}{12} = 7.333 \quad \text{C}_4\text{H}_8\text{O}_2 \\ (2) \quad & 9.1 \left(\frac{1}{1} \right) = 9.1 \quad \frac{88}{1} = 88 \quad 2(\text{C}_2) 2(\text{H}_4) 2(\text{O}) \\ (3) \quad & 36.4 \left(\frac{1}{16} \right) = 2.275 \quad \frac{88}{16} = 5.5 \quad \text{C}_4\text{H}_8\text{O}_2 \end{aligned}$$



$7 - 1.3$

Determining Molecular Formulas True Formulas Answers

Ray Dexter



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