

# Reading Comprehension

Grades 8-12

NAME _____	DATE _____
<h2 style="margin: 0;">PROPERTIES OF WATER</h2>	
<h3 style="margin: 0;">Heat Capacity and High Heat of Vaporization</h3> <p>(1) <b>Heat capacity</b> refers to the ability of a substance to absorb heat. Water has a high capacity to absorb heat. It is so good at this, in fact, that the temperature of 1 cm<sup>3</sup> of water (one degree Celsius) takes 4.18 joules to raise its temperature by 1°C. This means that water can absorb a lot of heat without a significant increase in temperature. This property is vital for regulating body temperature in organisms. It allows organisms to absorb heat from the environment without changing atmospheric temperatures. Additionally, water's capacity helps to moderate temperature changes on the planet, preventing drastic fluctuations that could be harmful to life. It is essential for regulating Earth's climate, as it reflects large areas, such as oceans and lakes, is almost large around the environment and only small changes in water temperature affect the climate.</p> <p>(2) <b>Heat of vaporization</b> is the amount of energy required to change a liquid into a gas. Water has a high heat of vaporization, meaning it takes a lot of energy to turn water into steam. This property is crucial for the cooling process in organisms. For example, when a liquid evaporates, it takes heat away from the surface it is evaporating from. In the human body, this is how sweating works. Sweat glands release sweat onto the skin, and as the sweat evaporates, it takes heat away from the body, cooling it down. This is why we feel cooler when we are wet. In the environment, this property helps to regulate the temperature of the atmosphere. Water vapor in the air can absorb a lot of heat, and when it condenses, it releases that heat, helping to moderate the climate.</p>	
<h3 style="margin: 0;">Cohesion</h3> <p>(1) <b>Cohesion</b> is the attraction between water molecules. Water molecules are polar, meaning they have a partial positive charge on one side and a partial negative charge on the other. This causes them to be attracted to each other. This property is essential for the transport of water in plants. Water molecules can form hydrogen bonds with each other, creating a network of bonds that allows water to move up the plant. This is how water can travel from the roots to the leaves, even against the force of gravity. This property is also important for the structure of water. Water molecules are attracted to each other, which gives water a high surface tension. This is why water can form droplets and why it can support small insects on its surface.</p> <p>(2) <b>Capillary action</b> is the ability of water to move up a narrow tube. This is due to the combination of cohesion and adhesion. Cohesion is the attraction between water molecules, and adhesion is the attraction between water molecules and the walls of the tube. In plants, this property is essential for the transport of water. Water molecules are attracted to each other, and they are also attracted to the walls of the xylem vessels in the plant. This allows water to move up the plant, even against the force of gravity. This property is also important for the structure of water. Water molecules are attracted to each other, which gives water a high surface tension. This is why water can form droplets and why it can support small insects on its surface.</p>	
<h3 style="margin: 0;">Adhesion and Cohesion</h3> <p>(1) <b>Adhesion</b> is the attraction between water molecules and other substances. Water molecules are polar, meaning they have a partial positive charge on one side and a partial negative charge on the other. This causes them to be attracted to other substances. This property is essential for the transport of water in plants. Water molecules can form hydrogen bonds with the walls of the xylem vessels in the plant, which allows them to move up the plant. This is how water can travel from the roots to the leaves, even against the force of gravity. This property is also important for the structure of water. Water molecules are attracted to each other, which gives water a high surface tension. This is why water can form droplets and why it can support small insects on its surface.</p> <p>(2) <b>Capillary action</b> is the ability of water to move up a narrow tube. This is due to the combination of cohesion and adhesion. Cohesion is the attraction between water molecules, and adhesion is the attraction between water molecules and the walls of the tube. In plants, this property is essential for the transport of water. Water molecules are attracted to each other, and they are also attracted to the walls of the xylem vessels in the plant. This allows water to move up the plant, even against the force of gravity. This property is also important for the structure of water. Water molecules are attracted to each other, which gives water a high surface tension. This is why water can form droplets and why it can support small insects on its surface.</p>	
<h3 style="margin: 0;">Capillary Action in Plants</h3> <p>(1) <b>Capillary action</b> is the ability of water to move up a narrow tube. This is due to the combination of cohesion and adhesion. Cohesion is the attraction between water molecules, and adhesion is the attraction between water molecules and the walls of the tube. In plants, this property is essential for the transport of water. Water molecules are attracted to each other, and they are also attracted to the walls of the xylem vessels in the plant. This allows water to move up the plant, even against the force of gravity. This property is also important for the structure of water. Water molecules are attracted to each other, which gives water a high surface tension. This is why water can form droplets and why it can support small insects on its surface.</p> <p>(2) <b>Capillary action</b> is the ability of water to move up a narrow tube. This is due to the combination of cohesion and adhesion. Cohesion is the attraction between water molecules, and adhesion is the attraction between water molecules and the walls of the tube. In plants, this property is essential for the transport of water. Water molecules are attracted to each other, and they are also attracted to the walls of the xylem vessels in the plant. This allows water to move up the plant, even against the force of gravity. This property is also important for the structure of water. Water molecules are attracted to each other, which gives water a high surface tension. This is why water can form droplets and why it can support small insects on its surface.</p>	

**PROPERTIES**

**Adhesion and Cohesion:**

(a) Hydrogen bonds between water molecules give water the properties of adhesion and cohesion. **Adhesion** is the attraction between water molecules and other substances. **Cohesion** is the attraction between water molecules. Both are related to the fact that water is able to adhere to other materials, such as when dewdrops form on surfaces or when raindrops stick to a windshield. These hydrogen bonds between water molecules also give water a high degree of cohesion. **Cohesion** is the attraction between water molecules. It allows water molecules to stick together to form droplets.

**Diagram:**

The diagram shows two groups of water molecules. On the left, labeled "COHESION", three water molecules are shown with arrows pointing from the oxygen atom of one molecule to the hydrogen atom of another, representing hydrogen bonds between water molecules. On the right, labeled "ADHESION", a water molecule is shown with an arrow pointing from its oxygen atom to a green vertical line representing a plant cell wall, representing the attraction between water and another substance.

(b) Capillary action occurs due to the combined effects of adhesion and cohesion. **Capillary action** is the ability of water to move against gravity in narrow spaces. **Soil** is the common feature in plants in capillary action or the phenomenon. **Adhesion** forces water molecules to adhere to the walls of narrow spaces and cohesion forces **soil** other water molecules along, allowing water to move upwards against the force of gravity. In plants, capillary action allows water to move upward from the roots to the leaves, providing essential nutrients and water to the entire plant. This process is vital for the survival of plants.

**Diagram:**

The diagram illustrates capillary action in a plant. On the left, a flower is shown with a label "Capillary action in plants" pointing to its stem. On the right, a circular inset shows a cross-section of a plant stem. The stem is divided into two parts: the outer part is labeled "xylem" and the inner part is labeled "phloem". Arrows indicate the upward movement of water through the xylem.

## PROPERTIES OF WATER

**Hydrogen Bonds**

Hydrogen bonds are the attractive forces between water molecules. They are responsible for many of water's unique properties, such as its high boiling point and surface tension.

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# Directed Reading Properties Of Water

**Michael Brown**



## **Directed Reading Properties Of Water:**

This book delves into Directed Reading Properties Of Water. Directed Reading Properties Of Water is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Directed Reading Properties Of Water, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
  - Chapter 1: Introduction to Directed Reading Properties Of Water
  - Chapter 2: Essential Elements of Directed Reading Properties Of Water
  - Chapter 3: Directed Reading Properties Of Water in Everyday Life
  - Chapter 4: Directed Reading Properties Of Water in Specific Contexts
  - Chapter 5: Conclusion
2. In chapter 1, this book will provide an overview of Directed Reading Properties Of Water. This chapter will explore what Directed Reading Properties Of Water is, why Directed Reading Properties Of Water is vital, and how to effectively learn about Directed Reading Properties Of Water.
3. In chapter 2, the author will delve into the foundational concepts of Directed Reading Properties Of Water. This chapter will elucidate the essential principles that must be understood to grasp Directed Reading Properties Of Water in its entirety.
4. In chapter 3, the author will examine the practical applications of Directed Reading Properties Of Water in daily life. This chapter will showcase real-world examples of how Directed Reading Properties Of Water can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Directed Reading Properties Of Water in specific contexts. This chapter will explore how Directed Reading Properties Of Water is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Directed Reading Properties Of Water. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Directed Reading Properties Of Water.

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