

**Student Reading - The Unique Properties of Water**

*(The ensuing discussion is adapted from Campbell, P.L.A. 1996. Biology (4th edition). Benjamin/Cummings Publ. Co. Inc. Menlo Park, CA, USA.)*

**The Unique Structure of Water**

Polarity of water molecules results in hydrogen bonding. The water molecule is relatively simple in structure. Two hydrogen atoms are joined to a single oxygen atom by single covalent bonds.

Oxygen is more electronegative than the hydrogen atoms, which allows the electrons of the polar bonds to spend more time closer to the oxygen side of the molecule. The oxygen side becomes more negative in charge, and the hydrogen atoms have a slight positive charge. This forms the polar molecule.

The water molecule is shaped like an isosceles triangle, with a slight bond angle of 104.5 degrees at the oxygen nucleus. The weak Coulombic characteristics of the bonding of hydrogen atoms to the weakly electronegative oxygen atom result in both ionic and covalent states that simultaneously maintain the integrity of water. Water is one of the only compounds that possess these characteristics.

An electrostatic attraction occurs between the polar water molecules. The slight positive charged hydrogen atom is attracted to the slight negative charged oxygen atom of another water molecule. This weak attraction is called a hydrogen bond. Every water molecule is hydrogen bonded to its four nearest neighbors.

**Cohesion of Water Molecules**

When water is in liquid form, its weak hydrogen bonds are about one-twentieth as strong as a covalent bond. Hydrogen bonds constantly form and break. Each hydrogen bond lasts for a fraction of a second, but the molecules continuously form new bonds with other water molecules around them. At any time, a large percentage of water molecules are bonded to neighboring water molecules which gives water more structure than most other liquids. Collectively, the hydrogen bonds hold water together by the property of cohesion.

Cohesion due to hydrogen bonding contributes to the formation of waves and other water movements that occur in lakes. Water movements are integral components of the lake system and play an important role in the distribution of temperature, dissolved gases, and nutrients. These movements also determine the distribution of macroorganisms and plankton.

Related to cohesion is surface tension, a measure of how difficult it is to stretch or break the surface of a liquid. Water has a greater surface tension than all other liquids except mercury. At the interface between water and air is an ordered arrangement of water molecules which are hydrogen bonded to one another and the water below. The result is an interface surface or film

# Directed Reading Section Properties Of Water Answers

**JL Elias**



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