

# MASTE 2

## Synthesis & Decomposition Reactions

Part 1

Name: 40 points

Per: \_\_\_\_\_

A. Below the word equation, write the Chemical equation for the reaction.

B. Determine if the reaction is a Synthesis or Decomposition reaction.

1. <sup>metallic</sup> Copper reacts with <sup>diatomic</sup> Chlorine gas to produce <sup>ionic</sup> Copper (II) chloride.



+, Synthesis

2. Ammonia ( $\text{NH}_3$ ) reacts with Hydrochloric acid ( $\text{HCl}$ ) to produce <sup>ionic</sup> Ammonium chloride.



+, Synthesis

3. Aluminum Hydroxide produces Aluminum oxide and Water.



+, Decomposition

4. <sup>metallic</sup> Sodium metal reacts with <sup>diatomic</sup> Oxygen gas to produce <sup>ionic</sup> Sodium oxide.



+, Synthesis

5. <sup>ionic</sup> Copper (II) oxide produces <sup>metallic</sup> Copper metal and <sup>diatomic</sup> Oxygen gas.



-, Decomposition

# Decomposition And Synthesis Reactions Answer Key

**Oswaal Editorial Board**



## **Decomposition And Synthesis Reactions Answer Key:**

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