

CHAPTER 4 The Structure Of The Atom

1. Use the periodic table to complete the following table.

Element	Atomic Number	Protons	Electrons
a. Li			
b.			87
c.	93		
d. Hg			80
e.	81		
f.	75		
g. B			

2. Give the number of protons, electrons, and neutrons in each of the following atoms.

- $^{108}_{47}\text{Ag}$
- $^{40}_{20}\text{Ca}$
- $^{23}_{11}\text{Na}$

3. Name each isotope, and write it in symbolic notation.

- atomic number 26; mass number 56
- atomic number 29; mass number 64
- atomic number 17; mass number 37

4. How many protons, electrons, and neutrons are in each of the following isotopes?

- uranium-235
- hydrogen-3
- silicon-29

5. How many neutrons does europium-151 have? What is the isotope's mass number?

6. How many more neutrons does thorium-230 have than protons? How many electrons does thorium-230 have?

7. Show that the mass number and the number of protons are conserved in the following nuclear equation: $^{234}_{90}\text{U} \rightarrow ^{230}_{90}\text{Th} + ^4_2\text{He}$.

8. Give the mass number of each isotope.

- Be with 5 neutrons
- Ca with 39 neutrons
- Si with 16 neutrons
- Ti with 26 neutrons

9. Give the atomic number of each isotope.

- magnesium-25
- bromine-79
- antimony-121

10. Neon has two isotopes: neon-10 and neon-12.

- Which isotope has the greater mass?
- Which has more neutrons?
- Which has more protons?
- Which has more electrons?

11. Use the table below to calculate the atomic mass of element X. Then use the periodic table to identify the element. Show all your work.

Isotope	Mass (amu)	Percent Abundance
^{16}X	15.995	99.762
^{17}X	16.999	0.038
^{18}X	17.999	0.20

Defining The Atomic Chapter 4 Packet Answers

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Defining The Atomic Chapter 4 Packet Answers:

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