

The Determination of a Chemical Formula

John Dalton was an Englishman, a teacher, and an exceptional theoretical chemist. He developed and wrote the modern atomic theory at the turn of the 19th century (documents point to 1803). He was influenced by the experiments of two Frenchmen, Antoine Lavoisier and Joseph Louis Proust. A fundamental component of the modern atomic theory is that the mole ratios of elements in a compound will be small whole numbers (*law of definite proportions*). The whole number mole ratio is commonly referred to as the empirical formula of a compound.

One of the challenges in finding the proper chemical formula for a compound is that there may be more than one plausible mole ratio for the elements in that compound. Dalton called this the *law of multiple proportions*. For example, if you were testing a compound that contained iron and sulfur, the plausible chemical formula could be FeS or Fe₂S₃. However, if you determine the mass of iron and the mass of sulfur present in a given mass of the compound, you will be able to establish the true chemical formula of the compound.

In this experiment, you will use the law of definite proportions to find the chemical formula for a hydrated compound containing copper, chlorine, and water molecules locked in the crystal structure of the solid compound. The general formula for the compound is Cu_xCl_y•zH₂O, and its name is copper chloride hydrate. The variables: x, y, and z represent integers that will establish the proper chemical formula for this substance.

First, you will gently heat a sample of the compound to drive off the water of hydration. By measuring the mass of the sample before and after heating you can determine the amount of water in the sample and identify the value of z. (Recall Experiment 1)

Second, you will conduct a chemical reaction with the dried sample, which will produce elemental copper. By measuring the mass of copper that forms, you will have the necessary information to determine the moles of copper and chlorine in your sample, and you will be able to establish the proper chemical formula.

Objectives

In this experiment, you will:

- Determine the water of hydration in a copper chloride hydrate sample.
- Conduct a reaction between a solution of copper chloride and solid aluminum.
- Use the results of the reaction to determine the mass and moles of Cu and Cl in the reaction.
- Calculate the empirical formula of the copper chloride compound.

Materials

Crucible with cover

Unknown solid copper chloride hydrate

Determination Of A Chemical Formula Lab Answers



G Orfield

Determination Of A Chemical Formula Lab Answers:

Million Dollar Data: Building Confidence – Vol.1 Stephen DeMeo, 2020-07-31 Global warming our current and greatest challenge is without precedent Among the many consequences that are impacting our society one unanticipated concern involves scientific truth When the President of the United States and others in his administration declare that global warming is fake science it calls into question what real science is and what real school science should be I will argue that real science is quality science one that is based on the rigorous collection of reliable and valid data To collect quality data requires bending over backwards to get things right and this is exactly what makes science so special Truth is made when scientists go this extra yard and devise controlled experiments collect large data sets confirm the data and rationally analyze their results Making scientific truth sounds difficult to do in the science laboratory but in reality there are many straightforward ways that truth can be constructed In the first of two volumes I discuss twelve such ways I call them Confidence Indicators that can allow students to strongly believe in their data and their subsequent results Many of these methods are intuitive and can be used by young students on the late elementary level all the way up to those taking introductory college science courses As in life science is not without doubt In the second volume I introduce the concept of scientific uncertainty and the indicators used to calculate its magnitude I will show that science is about connecting confidence with uncertainty in a specific manner what I refer to as the Confidence Uncertainty Continuum expression This important relationship epitomizes the scientific enterprise as a search for probabilistic rather than absolute truth This two volume set will contain a variety of ways that data quality can be instituted into a science curriculum To support its use many of the examples that I will present involve science teachers as well as student work and feedback from different grade levels and in different scientific disciplines Specific chapters will be devoted to reviewing the academic literature on data quality as well as describing my own personal research on this important but often neglected topic

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