

**DE LA SALLE SANTIAGO ZOBEL SCHOOL - VERMOSA**  
Senior High School Department

UNIT: <u>3</u>	AY: <u>2022-2023</u>	TERM: <u>2</u>	SUBJECT: <u>GENERAL CHEMISTRY 1</u>
MODULE NO.: <u>3</u>	TYPE: <input type="checkbox"/> PRE <input type="checkbox"/> SELF <input type="checkbox"/> FOR <input checked="" type="checkbox"/> SUM	SIMULATION EXPERIMENT NO: <u>1</u>	
TITLE OF THE ACTIVITY: <u>CHEMICAL REACTIONS &amp; MOLE CONCEPT &amp; STOICHIOMETRY</u>			SCORE: <table border="1" style="width: 50px; height: 20px; text-align: center;">26</table>
ASSESSMENT TOPICS: <u>Experiment #1: Determination of a Chemical Formula</u>			
NAME: _____			GR. & SEC. _____
TEACHER: <u>Ms. Malyn B. Soliman, Ms. Angelaid M. Vizcarra, and Mr. Carl Maurice P. Francisco</u>			DATE: _____

**General Instructions:**

- Read the handouts on Experiment no. 1 and write a pre-laboratory report about it. Use [Google docs](#) as your writing tool and make sure to save it in your drive. Use the following template for the file name: GR. & SEC. GRP NO. PRELAB EXPERIMENT NO.1 (example: 11A\_GRP 1\_PRELAB EXPERIMENT NO.1)
- Assign **one** (1) member to upload the file in the [General Chemistry 1 Canvas Course](#) on or before the deadline. Follow your teacher's instructions on the date and time of the submission of your output.
- The pre-lab report must include the following parts:
  - Objectives**  
-Formulate three (3) objectives that are related to the experiment.
  - Key Concepts**  
-Write at least three (3) concepts on how to determine the chemical formula of a compound.
  - Materials**  
- In a tabular form, write all the materials/equipment needed to perform the experiment.
  - Safety Measures**  
-Write at least three (3) safety measures that you should manifest while conducting the experiment.
  - Procedure (flow chart)/Schematic diagram**  
-Use phrases in doing the schematic diagram.
  - References**  
(Use APA format in writing your references.)
- Read the rubrics (pre-lab report and peer evaluation) and attached sample format below before doing your pre-lab report.
- Answer the Peer evaluation form uploaded in the General Chemistry 1 Canvas Course after submitting your task.

**Introduction:**

When atoms of one element combine with those of another, the combining ratio is typically an integer or a simple fraction; 1:2, 1:1, and 2:3 are ratios one might encounter. The simplest formula of a compound expresses that atom ratio. Some substances with the ratios we listed include  $\text{CaCl}_2$ ,  $\text{KBr}$ ,  $\text{Ag}_2\text{O}$ , and  $\text{Fe}_2\text{O}_3$ . When more than two elements are present in a compound, the formula still indicates the atom ratio. Thus the substance with the formula  $\text{Na}_2\text{S}_2\text{O}_4$  indicates that the sodium, sulfur, and oxygen atoms occur in that compound in the ratio 2:1:4. Many compounds have more complex formulas than those we have noted, but the same principle applies.

To find the formula of a compound, we need to find the mass of each of the elements in a weighed sample of that compound. For example, if we resolved a sample of the compound  $\text{NaOH}$  weighing 40 grams into its elements, we would find that we obtained just about 23 grams of sodium, 16 grams of oxygen, and 1 gram of hydrogen. Since the atomic mass scale tells us that sodium atoms have a relative mass of 23, oxygen atoms have a relative mass of 16, and hydrogen atoms a relative mass of just about 1, we would conclude that the sample of  $\text{NaOH}$  contained equal numbers of Na, O, and H atoms. Since that is the case, the atom ratio Na:O:H is 1:1:1, and so the simplest formula is  $\text{NaOH}$ . In terms of moles, we can say that one mole of  $\text{NaOH}$ , 40 grams, contains one mole of Na, 23 grams, one mole of O, 16 grams, and one mole of H, 1 gram, where we define the molar mass to be that mass in grams equal numerically to the sum of the atomic masses in an element or a compound. From this kind of argument, we can conclude that the atom ratio in a compound is equal to the mole ratio. We get the mole ratio from chemical analysis, and from that, the formula of the compound is determined.

In this experiment, we will use these principles to find the formula of the compound with the general formula  $\text{Cu}_x\text{Cl}_y \cdot z\text{H}_2\text{O}$ , where the x, y, and z are the integers that, when known, establish the formula of the compound. (In expressing the formula of a compound like this one, where water molecules remain intact within the compound, we retain the formula of  $\text{H}_2\text{O}$  in the formula of the compound.)

The compound we will study, which is called copper chloride hydrate, turns out to be ideal for one's first venture into formula determination. It is stable, can be obtained in pure form, has a characteristic blue-green color that changes as the compound is changed chemically, and is relatively easy to decompose into the elements and water. In the experiment, we will first drive out the water, which is called the water of hydration, from an accurately weighed sample of the compound. This occurs if we gently heat the sample to a little over  $100^\circ\text{C}$ . As the water is driven out, the color of the sample changes from blue-green to a tan-brown color similar to that of tobacco. The compound formed is **anhydrous ("no water")** copper chloride. If we subtract its mass from that of the hydrate, we can determine the mass of the water that was driven off, and, using the molar mass of water, find the moles of  $\text{H}_2\text{O}$  that were in

# Determination Of A Chemical Formula Answers

**Neil D. Jespersen, Alison Hyslop**



## **Determination Of A Chemical Formula Answers:**

*A Textbook of Organic Chemistry, 22e* Arun Bahl & B S Bahl, 2019 With an increased focus on fundamentals this new edition of *A Textbook of Organic Chemistry* continues to present the time tested functional group approach to the subject This examination oriented book breaks the intricacies of Organic Chemistry into easy to understand steps which gives the student the necessary foundation to build upon learn and understand Organic Chemistry in a way that is efficient as well as long lasting *Understanding General Chemistry* Atef Korchef, 2022-03-07 *Understanding General Chemistry* details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the properties of matter Written in an easy to understand format with helpful pedagogy to fuel learning the book features main objectives at the beginning of each chapter get smart sections and check your reading section at the end of each chapter The text is filled with examples and practices that illustrate the concepts at hand In addition a summary and extensive MCQs exercises and problems with the corresponding answers and explanations are readily available Additional features include Alerts students to common mistakes and explains in simple ways and clear applications how to avoid these mistakes Offers answers and comments alongside sample problems enabling students to self evaluate their skill level Includes powerful methods easy steps simple and accurate interpretations and engaging applications to help students understand complex principles Provides a bridge to more complex topics such as solid state chemistry organometallic chemistry chemistry of main group elements inorganic chemistry and physical chemistry This introductory textbook is ideal for chemistry courses for non science majors as well as health sciences and preparatory engineering students *General Chemistry* Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie s classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie s many previous works This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value ***Understanding Essential Chemistry*** Max Diem, 2025-05-09 Enables students to understand apply and retain key concepts in general chemistry *Understanding Essential Chemistry* offers a unique and approachable supplement to standard general chemistry textbooks designed specifically to aid students in mastering fundamental principles Drawing on extensive classroom experience chemistry professor Max Diem presents key concepts in an uninterrupted flow allowing students to follow a clear and straightforward path to comprehension With a logical algebraic framework the book is structured to build students confidence by breaking down complex topics into manageable pieces and encouraging critical thinking at every step Aimed at STEM majors this book includes checkpoints with example problems and final answers to reinforce concepts and promote independent problem solving skills By methodically emphasizing basic understanding this hands on guide gives students the tools to grasp the core chemistry principles necessary for success in their courses labs and future studies A must have survival guide to boost student confidence in the subject the text Presents

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*Chemistry* John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science

*Conceptual Chemistry Volume I For Class XI* S.K. Jain & Shailesh K. Jain, 1998 Conceptual Chemistry Volume I For Class XI

*Oswaal Karnataka SSLC Question Bank Class 10 Science Book Chapterwise & Topicwise (For 2024 Exam)* Oswaal Editorial Board, 2023-06-15 Description of the product Latest Board Examination Paper 2023 Held in April 2023 with Board Model Answer Strictly as per the Revised Textbook syllabus blueprint design of the question paper Latest Board specified typologies of questions for exam success Perfect answers with Board Scheme of Valuation Handwritten Topper s Answers for exam oriented preparation KTBS Textbook Questions fully solved Crisp revision with Revision notes and Mind maps Hybrid learning with best in class videos 2 Model Papers solved for Examination Practice 3 Online Model Papers

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## **Table of Contents Determination Of A Chemical Formula Answers**

1. Understanding the eBook Determination Of A Chemical Formula Answers
  - The Rise of Digital Reading Determination Of A Chemical Formula Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Determination Of A Chemical Formula Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in a Determination Of A Chemical Formula Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Determination Of A Chemical Formula Answers
  - Personalized Recommendations
  - Determination Of A Chemical Formula Answers User Reviews and Ratings



- Determination Of A Chemical Formula Answers and Bestseller Lists
- 5. Accessing Determination Of A Chemical Formula Answers Free and Paid eBooks
  - Determination Of A Chemical Formula Answers Public Domain eBooks
  - Determination Of A Chemical Formula Answers eBook Subscription Services
  - Determination Of A Chemical Formula Answers Budget-Friendly Options
- 6. Navigating Determination Of A Chemical Formula Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Determination Of A Chemical Formula Answers Compatibility with Devices
  - Determination Of A Chemical Formula Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Determination Of A Chemical Formula Answers
  - Highlighting and Note-Taking Determination Of A Chemical Formula Answers
  - Interactive Elements Determination Of A Chemical Formula Answers
- 8. Staying Engaged with Determination Of A Chemical Formula Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Determination Of A Chemical Formula Answers
- 9. Balancing eBooks and Physical Books Determination Of A Chemical Formula Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Determination Of A Chemical Formula Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Determination Of A Chemical Formula Answers
  - Setting Reading Goals Determination Of A Chemical Formula Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Determination Of A Chemical Formula Answers
  - Fact-Checking eBook Content of Determination Of A Chemical Formula Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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