

DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Name Max Lundy

Solve the problems below.

1. The empirical formula of a compound is NO_2 . Its molecular mass is 92 g/mol. What is its molecular formula?

$$\begin{aligned} (1) & \frac{14 + 2(16)}{46} = 2 \\ (2) & 92 \div 46 = 2 \\ (3) & 2(\text{N}) 2(\text{O}_2) \\ & \text{N}_2\text{O}_4 \end{aligned}$$

N_2O_4

2. The empirical formula of a compound is CH_2 . Its molecular mass is 70 g/mol. What is its molecular formula?

$$\begin{aligned} (1) & \frac{12 + 2}{14} = 5 \\ (2) & 70 \div 14 = 5 \\ (3) & 5(\text{C}) 5(\text{H}_2) \\ & \text{C}_5\text{H}_{10} \end{aligned}$$

C_5H_{10}

3. A compound is found to be 40.0% carbon, 6.7% hydrogen and 53.3% oxygen. Its molecular mass is 60 g/mol. What is its molecular formula?

$$\begin{aligned} (1) & 40 \left(\frac{1}{12} \right) = 3.333 \quad \frac{60}{12} = 5 \quad \text{C}_5\text{H}_{10}\text{O} \\ (2) & 6.7 \left(\frac{1}{1} \right) = 6.7 \quad \frac{60}{1} = 60 \quad 2(\text{C}) 2(\text{H}_6) 2(\text{O}) \\ (3) & 53.3 \left(\frac{1}{16} \right) = 3.333 \quad \frac{60}{16} = 3.75 \quad \text{C}_2\text{H}_4\text{O}_2 \end{aligned}$$

$\text{C}_2\text{H}_4\text{O}_2$

4. A compound is 64.9% carbon, 13.5% hydrogen and 21.6% oxygen. Its molecular mass is 74 g/mol. What is its molecular formula?

$$\begin{aligned} (1) & 64.9 \left(\frac{1}{12} \right) = 5.4083 \quad \frac{74}{12} = 6.166 \quad \text{C}_4\text{H}_{10}\text{O} \\ (2) & 13.5 \left(\frac{1}{1} \right) = 13.5 \quad \frac{74}{1} = 74 \quad 4(12) + 10 + 16 = 74 \\ (3) & 21.6 \left(\frac{1}{16} \right) = 1.35 \quad \frac{74}{16} = 4.625 \quad \text{C}_4\text{H}_{10}\text{O} \end{aligned}$$

$\text{C}_4\text{H}_{10}\text{O}$

5. A compound is 54.5% carbon, 9.1% hydrogen and 36.4% oxygen. Its molecular mass is 88 g/mol. What is its molecular formula?

$$\begin{aligned} (1) & 54.5 \left(\frac{1}{12} \right) = 4.5416 \quad \frac{88}{12} = 7.333 \quad \text{C}_3\text{H}_4\text{O} \\ (2) & 9.1 \left(\frac{1}{1} \right) = 9.1 \quad \frac{88}{1} = 88 \quad 3(12) + 4 + 16 = 56 \\ (3) & 36.4 \left(\frac{1}{16} \right) = 2.275 \quad \frac{88}{16} = 5.5 \quad 2(\text{C}_3) 2(\text{H}_4) 2(\text{O}) \\ & \text{C}_6\text{H}_8\text{O}_2 \end{aligned}$$

$\text{C}_6\text{H}_8\text{O}_2$

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Determining Molecular Formula True Formulas

Anthony S. Fauci



Determining Molecular Formula True Formulas:

Determining Molecular Formula True Formulas Book Review: Unveiling the Power of Words

In some sort of driven by information and connectivity, the energy of words has be evident than ever. They have the capacity to inspire, provoke, and ignite change. Such could be the essence of the book **Determining Molecular Formula True Formulas**, a literary masterpiece that delves deep to the significance of words and their affect our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall affect readers.

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