# Worksheet 10 Determining Empirical Formulas

What is the empirical formula (lowest whole number ratio) of the compounds below Compound that has 60.g carbon and 20. g hydrogen 2. Compound that has 26.35 g potassium and 23.65 g chloride 3. Compound has 22.1 g aluminum, 25.4 g phosphorus and 52.5 g oxygen 4. Compound is 13% magnesium and 87% bromine 5. Compound is 32.4 % sodium, 22.5% sulfur, and 45.1% oxygen 6. The empirical formula of a compound is NO2. Its molecular mass is 92 g/mol. What is its molecular formula 7. A compound is found to be 40.0% carbon, 6.7% hydrogen and 53.5% oxygen. Its molecular mass is 70 g/mol. What is the compounds molecular formula?

# **Determining Empirical Formulas Answer Key 75 Carbon**

Ji Su Park, David Camacho, Stefanos Gritzalis, James J. Park

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