

Test 4 bonding and VSEPR v3

Name key
 Date _____

True/False

(Circle T or F to indicate whether the statement is true or false)

- 1 = ☐ 1. Covalent bonding occurs when a metal reacts with a nonmetal.
- 3 = ☐ 2. The greater the difference in electronegativity between two bonded atoms, the more polar the bond.
- 1 = ☐ 3. Ionic bonding occurs between atoms with small differences in electronegativity.
- 1 = ☐ 4. The hydrogen bond can be formed as long as the molecule contains hydrogen.
- 3 = ☐ 5. N_2 is an example of a covalent bond.
- 1 = ☐ 6. CH_4 has ionic bonds.
- 3 = ☐ 7. A bond in which one atom donates both electrons of a pair is a coordinate covalent bond.

Multiple Choice Write the letter of the choice that best completes the statement or answers the question on the line.

- a 8. Which of the following formulas is incorrect?
 a. AlO_3 b. KBr c. KNO_3 d. Li_2O e. $CaCl_2$
- d 9. Which of the following bonds would be the least polar yet still be considered polar covalent?
 a. $Mg-O$ b. $C-O$ c. $O-O$ d. $H-O$
- b 10. Consider the drawings below



Which of the following statements are true?

- I. The electrons in each molecule tend to be attracted to the most electronegative element.
 - II. Each molecular drawing follows the electron sea model.
 - III. Both HF and CO_2 are linear molecules and nonpolar.
 - IV. The bond angles of H_2O are slightly less than 109.5° because the lone pair compresses the angles between the bonding pairs.
- a. I, III
 b. I, IV
 c. I, II, III
 d. II, IV
 e. All of the above statements (I - IV) are correct.

- a 12. Which molecule would have the strongest tendency to form hydrogen bonds with other chemical molecules?
 a. NH_3 b. CH_4 c. H_2 d. HBr e. C_2H_6

- c 13. You may need to draw the Lewis structures for the following compounds to assist you in answering this question.



How many of the compounds are nonpolar?

- a. 1 b. 2 c. 3 d. 4 e. 5

- a 14. You may need to draw the Lewis structures for the following compounds to assist you in answering this question. (Same compounds as previous question)



Which compound has bond angles of 109.5° around the central atom?

- a. CH_4 b. BH_3 c. $XeCl_4$ d. SF_6 e. HCl

- b 15. Which of the following compounds contains an ionic bond?

- a. $HCl(g)$ b. $AgCl$ c. CCl_4 d. SO_2 e. O_2

- a 16. Which of the following best describes all the intermolecular forces exhibited by a pure sample of PH_3 ?
 a. dispersion only b. dipole-dipole and hydrogen bonding c. dispersion and hydrogen bonding
 d. dispersion, dipole-dipole, and hydrogen bonding e. dispersion and dipole-dipole

- a 17. Which of the following compounds contains one or more covalent bonds?

- a. Cl_2 b. $LiCl$ c. Na_2O d. $CoCl_2$ e. MgS

- c 18. The carbonate ion has the formula CO_3^{2-} . What is the correct formula for sodium carbonate?

- a. $NaCO_3$ b. Na_2CO_3 c. $NaCO_2$ d. Na_2CO_2 e. $NaCO_4$

- c 19. If atom X forms a diatomic molecule with itself, the bond is

- a. ionic b. polar covalent c. nonpolar covalent d. polar-covalent covalent

Complete: 1 point each

10. Phosphorus(V) oxide has the formula P_2O_5

11. The correct name for PO_4^{3-} is phosphate



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M Walker



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