#### Honors Chemistry I Reaction Prediction Worksheet

For the following reactants: (1) tell the type of reaction; (2) predict the products (if no reaction, write NR); (3) write the net ionic equation for all ppt, reactions. Make sure to include the proper state of matter for all rectants and products. If no state is indicated for compounds, they are assumed to be in aqueous solution (aq).

1. mixing solutions of lead nitrate and potassium chloride

DR 
$$Pb^{+2}_{(aq)} + 2 Cl^{-1}_{(aq)} \rightarrow PbCl_{2(s)}$$

2. dropping solid magnesium metal into hydrochloric acid

$$SR = Mg_{(n)} + 2 H^{+1}_{(nq)} \rightarrow Mg^{+2}_{(nq)} + H_{2(g)}$$

3. adding solid tin to sulfuric acid

$$SR = Sn_{(x)} + 2 H^{+1}_{(aq)} \rightarrow Sn^{+2}_{(aq)} + H_{2(g)}$$

4. bubbling chlorine gas into hydrobromic acid

$$SR = Cl_{2(g)} + 2 Br^{-1}_{(aq)} \rightarrow 2 Cl^{-1}_{(aq)} + Br_{2(liq)}$$

5. passing sulfur dioxide gas over solid calcium oxide

SYN 
$$SO_{2(g)} + CaO_{(i)} \rightarrow CaSO_{3(i)}$$

6. putting solid lithium oxide into water

SYN 
$$\text{Li}_2O_{(s)} + \text{H}_2O_{(liq)} \rightarrow 2 \text{Li}^{+1}_{(aq)} + 2 \text{OH}^{-1}_{(aq)} \text{ or } 2 \text{LiOH}_{(aq)}$$

7. heating silver oxide

DEC 
$$2 \text{ Ag}_2\text{O}_{(0)} \rightarrow 4 \text{ Ag}_{(1)} + \text{O}_{2(g)}$$

8. heating sodium oxide

heating strontium sulfate

10. combining barium with iodine and heating

SYN 
$$Ba_{(s)} + I_{2(s)} \rightarrow BaI_{2(s)}$$

# **Honors Chemistry Reaction Prediction Packet**

**Léo Pomar** 

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to new geochemical behavior that might not be understood or predicted by considering smaller or larger scales alone

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