

CHAPTER 8 REVIEW

Chemical Equations and Reactions

SHORT ANSWER Answer the following questions in the space provided.

2. _____ According to the law of conservation of mass, the total mass of the reacting substances is
- (a) always more than the total mass of the products.
 - (b) always less than the total mass of the products.
 - (c) sometimes more and sometimes less than the total mass of the products.
 - (d) always equal to the total mass of the products.

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 2. In the reaction described by the word equation
- sodium + water \rightarrow sodium hydroxide + hydrogen gas
- a. sodium hydroxide is a product.
 - b. sodium is a product.
 - c. water is a reactant.
 - d. Both (a) and (c).
- _____ 3. Which of the following indicates that a chemical equation is balanced?
- a. The numbers of atoms of each element are the same on both sides of the equation.
 - b. All of the coefficients are the same.
 - c. The numbers of molecules on each side of the equation are equal.
 - d. The sums of the coefficients on each side of the equation are equal.
- _____ 4. From a complete and correctly written chemical equation, you can obtain the
- a. chemical formulas of the reactants and products.
 - b. relative amounts of the reactants and products.
 - c. physical states of the reactants and products.
 - d. All of the above.
- _____ 5. To balance a chemical equation, you must
- a. adjust the subscripts.
 - b. adjust the number of products.
 - c. adjust the number of reactants.
 - d. adjust the coefficients.
- _____ 6. The formation of a solid product in a chemical reaction is represented by the symbol
- a. (aq).
 - b. (g).
 - c. (l).
 - d. (s).
- _____ 8. In the expression 3CO_2 , the numbers 3 and 2 are, respectively,
- a. a subscript and a coefficient.
 - b. a coefficient and a subscript.
 - c. two subscripts.
 - d. two coefficients.
- _____ 9. What is the total number of molecules that is represented by the following chemical equation? (Add up the coefficients.)
- $\text{NH}_4\text{NO}_2 \rightarrow \text{N}_2 + 2\text{H}_2\text{O}$
- a. 3
 - b. 4
 - c. 8
 - d. 16
- _____ 10. Which equation below violates the law of conservation of mass?
- a. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 - b. $\text{KCl} + \text{Br} \rightarrow \text{KBr} + \text{Cl}_2$
 - c. $2\text{Fe}_2\text{O}_3 + 3\text{C} \rightarrow 4\text{Fe} + 3\text{CO}_2$
 - d. $\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$
- _____ 11. Which of the following equations represents the balanced equation for the reaction of iron and oxygen?
- a. $2\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 - b. $\text{Fe} + 3\text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 - c. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 - d. $3\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
- _____ 12. Which coefficients correctly balance the formula equation below?
- $\text{KClO}_3(s) \rightarrow \text{KCl}(s) + \text{O}_2(g)$
- a. 1, 1, 1
 - b. 1, 1, 3
 - c. 2, 2, 3
 - d. 2, 1, 1
- _____ 13. Which equation is *not* balanced?
- a. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 - b. $4\text{H}_2 + 2\text{O}_2 \rightarrow 4\text{H}_2\text{O}$
 - c. $\text{H}_2 + \text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$
 - d. $2\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
- _____ 14. In which type of reaction do two or more compounds react to form one product?
- a. synthesis (AKA combination)
 - b. decomposition
 - c. combustion
 - d. single-replacement
- _____ 15. Which of the following is true regarding single-replacement reactions?
- a. Only one product is formed.
 - b. There is only one reagent.
 - c. It requires oxygen.
 - d. One element replaces another in a compound.

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