

# STD XII - CHEMISTRY

Time : 3 hrs.]

MARCH - 2015

[Marks : 70

**Note :** (i) All questions are compulsory.

(ii) Answer to the two sections should be written in the same answer book.

(iii) Figures to the right hand side indicate full marks.

(iv) Write balanced chemical equations and draw neat and labelled diagrams wherever necessary.

(v) Use of logarithmic table is allowed.

(vi) Answer to every new question must be started on a new page.

## SECTION - I

- Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :** [7]
- (i) p-type semi-conductors are made by mixing silicon with impurities of \_\_\_\_\_.  
(a) germanium (b) boron (c) arsenic (d) antimony
- (ii) Amongst the following identify the criterion for a process to be at equilibrium :  
(a)  $\Delta G < 0$  (b)  $\Delta G > 0$  (c)  $\Delta S_{\text{total}} = 0$  (d)  $\Delta S < 0$
- (iii) Colligative property depends only on \_\_\_\_\_ in a solution.  
(a) Number of solute particles. (b) Number of solvent particles.  
(c) Nature of solute particles. (d) Nature of solvent particles.
- (iv) The charge of how many coulomb is required to deposit 1.0 g of sodium metal (molar mass  $23.0 \text{ g mol}^{-1}$ ) from sodium ions is \_\_\_\_\_.  
(a) 2098 (b) 96500 (c) 193000 (d) 4196
- (v) What is the chemical composition of malachite ?  
(a)  $\text{CuO} \cdot \text{CuCO}_3$  (b)  $\text{Cu}(\text{OH})_2 \cdot \text{CuCO}_3$  (c)  $\text{CuO} \cdot \text{Cu}(\text{OH})_2$  (d)  $\text{Cu}_2\text{O} \cdot \text{Cu}(\text{OH})_2$
- (vi) The element that does NOT exhibit allotropy is \_\_\_\_\_.  
(a) As (b) Sb (c) Bi (d) N
- (vii) The integrated rate equation for first order reaction is  $A \xrightarrow{\quad} \text{products}$  \_\_\_\_\_.  
(a)  $K = 2.303t \log_{10} \frac{[A]_0}{[A]_t}$  (b)  $K = -\frac{1}{t} \ln \frac{[A]_t}{[A]_0}$   
(c)  $K = \frac{2.303}{t} \log_{10} \frac{[A]_t}{[A]_0}$  (d)  $K = \frac{1}{t} \ln \frac{[A]_t}{[A]_0}$

- Q. 2. Answer any SIX of the following :** [12]
- (i) Define the following terms :  
(a) Enthalpy of fusion  
(b) Enthalpy of atomization
- (ii) Derive van't Hoff general solution equation.
- (iii) Explain impurity defect in stainless steel with diagram.
- (iv) Derive the relation between half life and rate constant for a first order reaction.
- (v) Draw neat and labelled diagram of dry cell.
- (vi) Explain the structure of sulphur dioxide.
- (vii) What is calcination ? Explain it with reactions.
- (viii) Arrange the following reducing agents in the order of increasing strength under

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**JE Gale**



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