

### Observations:

1. Solution in a burette : known solution
2. Solution by a pipette : 0.1 M oxalic acid solution
3. Solution in conical flask : 0.1 M oxalic acid + Dil. H<sub>2</sub>SO<sub>4</sub>
4. Indicator : oxalic acid solution known
5. End point : colorless to light pink
6. Chemical Equation :  $2\text{KMnO}_4 + 5\text{H}_2\text{C}_2\text{O}_4 + 3\text{H}_2\text{SO}_4 \longrightarrow \text{K}_2\text{SO}_4 + 2\text{MnSO}_4 + 8\text{H}_2\text{O} + 10\text{CO}_2\uparrow$

Observation Table:

Burette level	Pilot reading	Burette reading in mL			C.B.R.
		I	II	III	
Final	9.5	9.5	9.5	9.5	9.5 (x) mL
Initial	to	0.0	0.0	0.0	
Difference	9.5 mL	9.5	9.5	9.5	

Calculation: From above chemical equation



$2 \times 158 \text{ g (316 g)} = 1000 \text{ mL } 5 \text{ M (molar mass of KMnO}_4 = 158 \text{ g/mol)}$

$\therefore 1000 \text{ mL } 5 \text{ M oxalic acid} = 316 \text{ g of KMnO}_4$

$$\therefore 10 \text{ mL } 0.1 \text{ M oxalic acid} = \frac{316 \times 10 \times 0.1}{1000 \times 5}$$

$$= 0.0632 \text{ g of KMnO}_4$$

$\therefore$  Hence (x) C.B.R. mL of KMnO<sub>4</sub> solution contains = 0.0632 g of KMnO<sub>4</sub>

$$\therefore 1000 \text{ mL of KMnO}_4 \text{ contains} = \frac{0.0632 \times 1000}{9.5(x) \text{ CBR}}$$

Hence molarity of KMnO<sub>4</sub> in the solution is  $\sim \frac{0.0632 \times 1000}{9.5(x) \text{ CBR} \times 158}$

$$= \frac{0.4}{9.5(x) \text{ CBR}}$$

$$= 0.04211 \dots \text{ M}$$

Result: The molarity of KMnO<sub>4</sub> solution is 0.04211... M

Remark and sign of teacher:

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Space for log calculation

Number	log
$\log [0.4 \div 9.5]$	
1.6021	
- 0.9777	
<u>Ans: 2.6244</u>	

$$= 4.21 \times 10^{-2}$$

$$= 0.04211$$

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**Jim Stamell**



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### **Table of Contents Hsc Chemistry Part1 Target Solution Pdf**

1. Understanding the eBook Hsc Chemistry Part1 Target Solution Pdf
  - The Rise of Digital Reading Hsc Chemistry Part1 Target Solution Pdf
  - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Chemistry Part1 Target Solution Pdf
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hsc Chemistry Part1 Target Solution Pdf
  - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Chemistry Part1 Target Solution Pdf
  - Personalized Recommendations

- Hsc Chemistry Part1 Target Solution Pdf User Reviews and Ratings
- Hsc Chemistry Part1 Target Solution Pdf and Bestseller Lists
- 5. Accessing Hsc Chemistry Part1 Target Solution Pdf Free and Paid eBooks
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  - Hsc Chemistry Part1 Target Solution Pdf Budget-Friendly Options
- 6. Navigating Hsc Chemistry Part1 Target Solution Pdf eBook Formats
  - ePub, PDF, MOBI, and More
  - Hsc Chemistry Part1 Target Solution Pdf Compatibility with Devices
  - Hsc Chemistry Part1 Target Solution Pdf Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Hsc Chemistry Part1 Target Solution Pdf
  - Highlighting and Note-Taking Hsc Chemistry Part1 Target Solution Pdf
  - Interactive Elements Hsc Chemistry Part1 Target Solution Pdf
- 8. Staying Engaged with Hsc Chemistry Part1 Target Solution Pdf
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Hsc Chemistry Part1 Target Solution Pdf
- 9. Balancing eBooks and Physical Books Hsc Chemistry Part1 Target Solution Pdf
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Hsc Chemistry Part1 Target Solution Pdf
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Chemistry Part1 Target Solution Pdf
  - Setting Reading Goals Hsc Chemistry Part1 Target Solution Pdf
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Chemistry Part1 Target Solution Pdf
  - Fact-Checking eBook Content of Hsc Chemistry Part1 Target Solution Pdf

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
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