

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :

[7]

- (i) The Arrhenius equation is _____.
- (a) $k = A e^{\frac{RT}{E_a}}$ (b) $A = k e^{\frac{-E_a}{RT}}$
- (c) $k = A e^{\frac{-RT}{E_a}}$ (d) $k = A e^{\frac{-K_a}{RT}}$
- (ii) If the enthalpy of vaporisation of water at 100°C is 186.5 J . mol⁻¹, the entropy of vaporization will be _____.
- (a) 4.0 J . K⁻¹. mol⁻¹ (b) 3.0 J . K⁻¹. mol⁻¹
- (c) 1.5 J . K⁻¹. mol⁻¹ (d) 0.5 J . K⁻¹. mol⁻¹
- (iii) The atomicity of sulphur in orthorhombic sulphur is _____.
- (a) 8 (b) 6 (c) 4 (d) 2
- (iv) The major binding force in diamond is _____.
- (a) Covalent bond (b) Ionic bond
- (c) Metallic bond (d) Co-ordinate covalent bond
- (v) The boiling point of water at high altitude is low, because _____.
- (a) the temperature is low. (b) the atmospheric pressure is low.
- (c) the temperature is high. (d) the atmospheric pressure is high.
- (vi) The molar conductivity of cation and anion of salt BA are 180 and 220 mhos respectively. The molar conductivity of salt BA at infinite dilution is _____.
- (a) 90 mhos.cm². (b) 110 mhos.cm².mol⁻¹
- (c) 200 mhos.cm².mol⁻¹ (d) 400 mhos.cm².mol⁻¹
- (vii) What is the process in which concentrated ore is reduced to the corresponding metal by heating at high temperature with a reducing agent?
- (a) Polling (b) Pyrometallurgy (c) Hydrometallurgy (d) Calcination

Q. 2. Answer any THREE of the following :

[9]

- (i) Describe anomalous behaviour of oxygen as compared with other elements of group 16 with reference to –
- (a) Magnetic property (b) Oxidation state (c) Hydrides
- (ii) What is the value of $\Delta S_{\text{surr.}}$ for the following reaction at 298 K –
- $$6\text{CO}_{2(g)} + 6\text{H}_2\text{O}_{(l)} \longrightarrow \text{C}_6\text{H}_{12}\text{O}_{6(s)} + 6\text{O}_{2(g)}$$
- Given that : $\Delta G^\circ = 2879 \text{ kJ mol}^{-1}$, $\Delta S = -210 \text{ JK}^{-1} \text{ mol}^{-1}$.
- (iii) Sucrose decomposes in acid solution to give glucose and fructose according to the first order rate law. The half life of the reaction is 3 hours. Calculate fraction of sucrose which will remain after 8 hours.
- (iv) A solution containing 0.73 g of camphor (molar mass 152 g . mol⁻¹) in 36.8 g of acetone (boiling point 56.3°C) boils at 56.55°C. A solution of 0.564 g of unknown compound in the same weight of acetone boils at 56.46°C. Calculate the molar mass of the unknown compound.

Hsc Chemistry Question 1st Paper 2014

John Michels (Journalist)



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Cambridge Checkpoints HSC Chemistry 2017-19 Maureen Slade,Roger Slade,2016-06-29 Earth's Early Atmosphere and Oceans, and The Origin of Life George H. Shaw,2015-10-07 This book provides a comprehensive treatment of the chemical nature of the Earth s early surface environment and how that led to the origin of life This includes a detailed discussion of the likely process by which life emerged using as much quantitative information as possible The emergence of life and the prior surface conditions of the Earth have implications for the evolution of Earth s surface environment over the following 2 2 5 billion years The last part of the book discusses how these changes took place and the evidence from the geologic record that supports this particular version of early and evolving conditions **Understanding Student Participation and Choice in Science and Technology Education** Ellen Karoline Henriksen,Justin Dillon,Jim Ryder,2014-09-17 Drawing on data generated by the EU s Interests and Recruitment in Science IRIS project this volume examines the issue of young people s participation in science technology engineering and mathematics education With an especial focus on female participation the chapters offer analysis deploying varied theoretical frameworks including sociology social psychology and gender studies The material also includes reviews of relevant research in science education and summaries of empirical data concerning student choices in STEM disciplines in five European countries Featuring both quantitative and qualitative analyses the book makes a substantial contribution to the developing theoretical agenda in STEM education It augments available empirical data and identifies strategies in policy making that could lead to improved participation and gender balance in STEM disciplines The majority of the chapter authors are IRIS project members with additional chapters written by specially invited contributors The book provides researchers and policy makers alike with a comprehensive and authoritative exploration of the core issues in STEM educational participation Science John Michels (Journalist),2008 *Success One Chemistry* 2018 Edit,Mark Dixon,2018-02-12 Excel Success One HSC Chemistry contains 2001 2003 and 2008 2016 past HSC questions with detailed answers written by experienced HSC markers a Topic Index a Mark Maximizer Guide and a Glossary of Key Verbs This title helps you get the results you want by practising actual HSC papers and answering HSC level questions Electronic Engineering ,1954 **Australian Education Index** ,1980 **Checkmate HSC Chemistry** Surendra Verma,1986 *2 Unit Chemistry HSC Course Revision* New South Wales. Board of Studies,1994 **Success One Physics** 2018 Edit,Mark Dixon,2018-02-12 Excel Success One HSC Physics contains 2001 2017 past HSC questions with detailed answers written by experienced HSC markers a Topic Index a Mark Maximizer Guide and a Glossary of Key Verbs This title helps you get the results you want by practising actual HSC papers and answering HSC level questions **Checkmate HSC Chemistry** Surendra Verma,1986-01-01 **Excel Success One HSC Physics** 2019 Edit,2019-02-12

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