পরীকারী কপুরা, এখানে বোর্র পরীকা, ক্যাডোট কলেলগর শীর্ষান্ধানীর কলেলগমূহের পরীক্ষা এবং মহেল টেন্টের ব্যাত্তর পূর্ণাল উত্তর অধ্যায়ভিত্তিক দেবছা ময়েছে। এগুলো অনুশীলন করনে কৃষি এ অধ্যায়টি লেকে সেকোনো স্থানশীল জনামূদ্

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- ক, মোলার দ্রবণ কীয়
- "কেলাস পানিত্র লকণের দ্রাকারা, তাপমারা বৃশ্বির সাথে সাথে এক পরীয়ে বৃশ্বি না পেয়ে য়াস পায়"— ব্যাব্যা করে।
- চবশ দৃটি মিশ্রিত করলে, থিপ্রদটি কী পরিমাণ অগ্নীয় বা ফারীয় মবে— বিশ্লেষণ করো।

#### ১ নং প্রয়ের উত্তর

1000 mL চৰণে কোন দ্ৰবের 1 mol দ্রবীভূত থাকলে উপ্ত দ্রবণকে
মোলার দ্রবন বলে।

ত্রলাস পানিত্র গবণের প্রাব্যান্তা তাপমাত্রা বৃশ্বির সাথে সাথে এক পর্যান্ত বৃশ্বির না পেরা ছাস পার। এর কারণ, তাপমাত্রা বৃশ্বির সাথে সাবে প্রব্যান্ত বৃশ্বির বিধান পার। এর কারণ, তাপমাত্রা বৃশ্বির সাথে সাবে প্রব্যান বৃশ্বির পোনে থাকার একটি নির্নিন্দ তাপমাত্রার কেলাস পানি কুর গরুণ নির্নান্ত হয়। তথান প্রাব্যান্ত হাস পোনে থাকে। উদাহরণমূল্প Na<sub>2</sub>SO<sub>4</sub>.10H<sub>2</sub>O এর কথা করা যায়। Na<sub>2</sub>SO<sub>4</sub>.10H<sub>2</sub>O এর পেরে পানিতে প্রবীত্ত হওয়ার সময় বিয়োজন প্রক্রিয়াটি ভাপহারী বিক্রিয়া কিন্তু 32.4°C তাপমাত্রার উপরে Na<sub>2</sub>SO<sub>4</sub>.10H<sub>2</sub>O যখন Na<sub>2</sub>SO<sub>4</sub> এ পরিণত হয় তথান বিয়োজন প্রক্রিয়াটি ভাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা কৃশ্বিজনিত কারণে প্রাব্যান্ত হস হটে।

#### 191

भ = SMV

-(0.5 × 98 × 0.25) g

= 12.25 g

অতএব, দ্বীভূত H₂SO₄ এর ভর ≒ 12.25 g

উদ্দীপকের A-পাত্রে অবস্থিত,

H<sub>2</sub>SO<sub>4</sub> এর আণবিক ভর,

M = 1 × 2 + 32 + 16 × 4

= 2 + 32 + 64

= 98 g/mol

রবণের ঘনমারা,

S = 0.5 M

রবণের আরতন,

V = 250 mL

= 250

1000 = 0.25L

ফরীভূত H<sub>2</sub>SO<sub>4</sub> এর ভর, w = 1

ভারত ন: H<sub>2</sub>SO<sub>4</sub> এর মোল সংখ্যা —  $\left(0.5 \times \frac{250}{1000}\right)$  med — 0.125 mod — 0.125 mod — 0.125 mod — 0.077 mod — 0.077 mod — 0.077 mod সংঘটিত বিভিন্ন: H<sub>2</sub>SO<sub>4</sub> + Na<sub>2</sub>CO<sub>3</sub> — 0.077 mod — 0.077

পাত্র-∧ এর প্রবংগর সাথে পাত্র-৪ এর প্রবংগর বিভিন্নার ফলে ১১৯ moi H₂SO₄ অবশিন্ট থাকায় মিশ্রিত প্রবংগর প্রকৃতি মধে অঞ্জীত

19.05

ett.

ভটা ১২ একজন ছাত্র লাবরেটরিতে X-পশ্বতিতে 10 mg বর জেছ হ পরিমাণ এবং Y-পশ্বতিতে 0.5 g পরিমাণ 2% ₩ HCl এর নমূর হ কাজ করে।

- ক, নোড কাকে বলে?
- খ. Cl<sub>2</sub> অপোলার- ব্যাখ্যা করে।।
- উদ্দীপকের এসিভটির মোলার ঘনমারা নির্ণয় করে।
- য়, উদ্দীপকের কোন পশ্বতি পরিবেশবান্ধর? বিশ্লেষণ করে। ২ নং প্রয়োর উত্তর

নিউরিয়াসের চারনিকে যে নির্দিউ ত্রিমাত্রিক স্থানে কোন ইত্রু পাওয়ার সম্ভাব্যতা শূন্য তাকে নোভ বলে।

সমযোজী বন্ধনে আবন্ধ উভয় পরমাণুর তড়িং কণাছকতা ক্র হলে অর্থাৎ তড়িং কণাব্যকতার পার্থকা শূন্য হলে সমযোজী ক বিশুন্ধ ও অপোলার হয়। আর তড়িং ঝণাব্যকতার পার্থকা 0.5 এর হ মর্লে অপোলার হয়। যেহেত CI এর তড়িং ঝণাব্যকতা 3.0 ৫ ট অপুতে তড়িং ঝণাব্যকতার পার্থকা (3 – 3)= 0, তাই CI<sub>2</sub> ক্রিন্ সমযোজী ও অপোলার যৌগ। উরোধ্য মৌলিক অপুর ক্রেন্তে এর্ণ হা থাকে।

ত্রী উদ্দীপকে বর্ণিত এগিচটি হলো 2%  $\left(\frac{W}{V}\right)$  HCI ।

2%  $\left(\frac{W}{V}\right)$  HCI এর অর্থ ফলো-

.. 1000 mL ... ... ...  $\frac{2 \times 1000}{100}$  g

# **Hsc First Year Chemistry 1st Paper Text**

**Michael Brown** 

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#### **Table of Contents Hsc First Year Chemistry 1st Paper Text**

- 1. Understanding the eBook Hsc First Year Chemistry 1st Paper Text
  - The Rise of Digital Reading Hsc First Year Chemistry 1st Paper Text
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Hsc First Year Chemistry 1st Paper Text
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hsc First Year Chemistry 1st Paper Text
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Hsc First Year Chemistry 1st Paper Text
  - Personalized Recommendations
  - Hsc First Year Chemistry 1st Paper Text User Reviews and Ratings
  - Hsc First Year Chemistry 1st Paper Text and Bestseller Lists
- 5. Accessing Hsc First Year Chemistry 1st Paper Text Free and Paid eBooks
  - Hsc First Year Chemistry 1st Paper Text Public Domain eBooks
  - Hsc First Year Chemistry 1st Paper Text eBook Subscription Services
  - Hsc First Year Chemistry 1st Paper Text Budget-Friendly Options
- 6. Navigating Hsc First Year Chemistry 1st Paper Text eBook Formats

- o ePub, PDF, MOBI, and More
- Hsc First Year Chemistry 1st Paper Text Compatibility with Devices
- Hsc First Year Chemistry 1st Paper Text Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Hsc First Year Chemistry 1st Paper Text
  - Highlighting and Note-Taking Hsc First Year Chemistry 1st Paper Text
  - Interactive Elements Hsc First Year Chemistry 1st Paper Text
- 8. Staying Engaged with Hsc First Year Chemistry 1st Paper Text
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Hsc First Year Chemistry 1st Paper Text
- 9. Balancing eBooks and Physical Books Hsc First Year Chemistry 1st Paper Text
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Hsc First Year Chemistry 1st Paper Text
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc First Year Chemistry 1st Paper Text
  - Setting Reading Goals Hsc First Year Chemistry 1st Paper Text
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc First Year Chemistry 1st Paper Text
  - Fact-Checking eBook Content of Hsc First Year Chemistry 1st Paper Text
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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