

## Organic Chemistry Laboratory

### #3 - Extraction

Extraction is a very important technique in organic chemistry. During the semester, and more so next semester, you will need to rely on extraction to separate a compound from a mixture of compounds. Primarily, you will do this to isolate and purify reaction products. Extraction can also be used to isolate compounds from plant or animal sources. Your success in this experiment is directly related to your understanding the pre-lab reading. Make sure you read this prior to the exercise and if you have questions try to get them answered.

#### Pre-lab Reading:

Read Chapters 15 in the Zubrick manual, which deal with the extraction process and the theory of extraction respectively. You should also read Chapter 10 in the Zubrick text that discusses drying agents.

#### Pre-lab assignment:

Sometimes a flowchart is helpful in understanding what is happening during a given process such as an extraction. The last page of this handout contains an incomplete flowchart for the separation of aniline, benzoic acid, and naphthalene. Complete this handout prior to the lab period. Also, for each of the following pairs of immiscible solvents indicate which would be on top and which would be on bottom, in a separatory funnel: **a)** water, ether; **b)** water, dichloromethane; **c)** 10% NaOH, hexane.

#### Procedure:

Obtain an unknown sample and weigh the container with the sample, record this weight. Dissolve your unknown in 50 mL of ether and transfer it in a 125 mL separatory funnel. (*Be sure that the stopcock is closed!*) You should rinse the original container with a small amount of ether and add this to the separatory funnel. Then weigh the empty, dry container and determine the mass of the unknown mixture.

To separate the basic component of your unknown extract the solution with three 15 mL portions of 5% aqueous HCl. Set up a ring stand to hold the separatory funnel while the layers separate. Combine all the aqueous layers into a 125 mL Erlenmeyer flask. Label this flask and set it aside for later.

Next, separate the acidic component of the mixture by extracting the remaining unknown solution with three 15 mL portions of 5% aqueous NaOH solution. Combine these aqueous extracts into a fresh 125 mL Erlenmeyer flask. Label this flask and set it aside.

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**JG Myers**



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