



Honda V65 Magna

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web start with 8 3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8 3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149 4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137 1 \text{ g } \text{H}_2\text{O}}{149 4 \text{ g } \text{H}_2\text{O}} \times 100 = 91 77\%$ problems 6 8 are similar to problem 5

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160 g
g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield
yield actual $\times 100$ 260.0 g $\times 100$ 86.01 theoretical 302.3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.6 8 only four molecules can be made 9 this amount cannot be
weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12
c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{text actual yield}}{\text{text theoretical yield}} \times 100$ percent
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 CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is
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