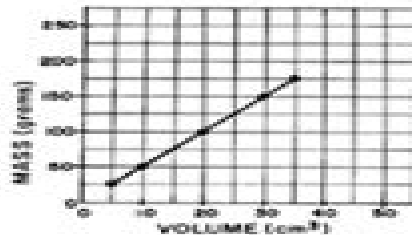


Honors Chemistry Written Final Exam Practice Test

1. Use the graph below to answer questions a-g:



a) What would be the mass of the substance if it has a volume of 15 cm³?

75 grams

b) What would be the volume of the substance if it has a mass of 175 g?

35 cm³

c) What is the density of the metal represented in the graph?

$$\text{Density} = \frac{\text{mass}}{\text{volume}} = \frac{75}{15} = 5.0 \text{ g/cm}^3$$

d) Calculate the mass of the substance above it takes up a volume 85 cm³.

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

$$5.0 = \frac{m}{85} \quad m = 425 \text{ g} = 430 \text{ g}$$

e) Calculate the volume of the substance above it has a mass of 325 grams.

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

$$5.0 = \frac{325}{V} \quad V = 65 \text{ cm}^3$$

f) Would this substance sink or float **in water** and why?

This substance would sink since it is more dense than water (water is 1 g/cm³)

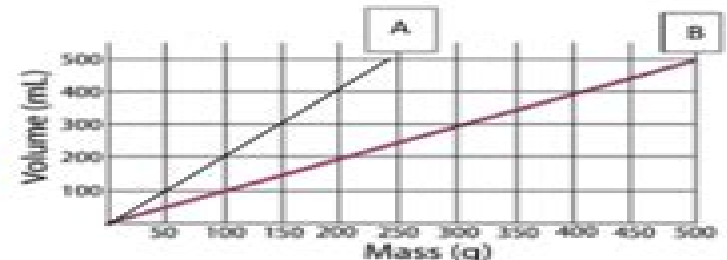
g) What does the slope of the line represent?

The density of the substance

For question h-i, use the graph to the right:

h) If the two substances were mixed together as liquids, which substance would sink to the bottom? How do you know?

Substance A is more dense since the slope of its line is steeper (you could have actually calculated density too) so substance A will sink.



i) Which would occupy a larger volume: 50 grams of substance A or 50 grams of substance B? Explain.

Since Substance B is has a lower density and you are comparing equivalent masses, Substance B would have to be taking up a larger volume.

Honors Chemistry Final Exam Mc Practice

AN Whitehead



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Table of Contents Honors Chemistry Final Exam Mc Practice

1. Understanding the eBook Honors Chemistry Final Exam Mc Practice
 - The Rise of Digital Reading Honors Chemistry Final Exam Mc Practice
 - Advantages of eBooks Over Traditional Books
2. Identifying Honors Chemistry Final Exam Mc Practice
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Honors Chemistry Final Exam Mc Practice
 - User-Friendly Interface
4. Exploring eBook Recommendations from Honors Chemistry Final Exam Mc Practice
 - Personalized Recommendations

- Honors Chemistry Final Exam Mc Practice User Reviews and Ratings
- Honors Chemistry Final Exam Mc Practice and Bestseller Lists
- 5. Accessing Honors Chemistry Final Exam Mc Practice Free and Paid eBooks
 - Honors Chemistry Final Exam Mc Practice Public Domain eBooks
 - Honors Chemistry Final Exam Mc Practice eBook Subscription Services
 - Honors Chemistry Final Exam Mc Practice Budget-Friendly Options
- 6. Navigating Honors Chemistry Final Exam Mc Practice eBook Formats
 - ePub, PDF, MOBI, and More
 - Honors Chemistry Final Exam Mc Practice Compatibility with Devices
 - Honors Chemistry Final Exam Mc Practice Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Honors Chemistry Final Exam Mc Practice
 - Highlighting and Note-Taking Honors Chemistry Final Exam Mc Practice
 - Interactive Elements Honors Chemistry Final Exam Mc Practice
- 8. Staying Engaged with Honors Chemistry Final Exam Mc Practice
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Honors Chemistry Final Exam Mc Practice
- 9. Balancing eBooks and Physical Books Honors Chemistry Final Exam Mc Practice
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Honors Chemistry Final Exam Mc Practice
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Honors Chemistry Final Exam Mc Practice
 - Setting Reading Goals Honors Chemistry Final Exam Mc Practice
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Honors Chemistry Final Exam Mc Practice
 - Fact-Checking eBook Content of Honors Chemistry Final Exam Mc Practice

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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