

**SECTION - I**

**Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :**

[7]

- (i) The Arrhenius equation is \_\_\_\_\_.
- (a)  $k = A e^{\frac{RT}{E_a}}$  (b)  $A = k e^{\frac{-E_a}{RT}}$
- (c)  $k = A e^{\frac{-RT}{E_a}}$  (d)  $k = A e^{\frac{-K_a}{RT}}$
- (ii) If the enthalpy of vaporisation of water at 100°C is 186.5 J . mol<sup>-1</sup>, the entropy of vaporization will be \_\_\_\_\_.
- (a) 4.0 J . K<sup>-1</sup>. mol<sup>-1</sup> (b) 3.0 J . K<sup>-1</sup>. mol<sup>-1</sup>
- (c) 1.5 J . K<sup>-1</sup>. mol<sup>-1</sup> (d) 0.5 J . K<sup>-1</sup>. mol<sup>-1</sup>
- (iii) The atomicity of sulphur in orthorhombic sulphur is \_\_\_\_\_.
- (a) 8 (b) 6 (c) 4 (d) 2
- (iv) The major binding force in diamond is \_\_\_\_\_.
- (a) Covalent bond (b) Ionic bond
- (c) Metallic bond (d) Co-ordinate covalent bond
- (v) The boiling point of water at high altitude is low, because \_\_\_\_\_.
- (a) the temperature is low. (b) the atmospheric pressure is low.
- (c) the temperature is high. (d) the atmospheric pressure is high.
- (vi) The molar conductivity of cation and anion of salt BA are 180 and 220 mhos respectively. The molar conductivity of salt BA at infinite dilution is \_\_\_\_\_.
- (a) 90 mhos.cm<sup>2</sup>. (b) 110 mhos.cm<sup>2</sup>.mol<sup>-1</sup>
- (c) 200 mhos.cm<sup>2</sup>.mol<sup>-1</sup> (d) 400 mhos.cm<sup>2</sup>.mol<sup>-1</sup>
- (vii) What is the process in which concentrated ore is reduced to the corresponding metal by heating at high temperature with a reducing agent?
- (a) Polling (b) Pyrometallurgy (c) Hydrometallurgy (d) Calcination

**Q. 2. Answer any THREE of the following :**

[9]

- (i) Describe anomalous behaviour of oxygen as compared with other elements of group 16 with reference to –
- (a) Magnetic property (b) Oxidation state (c) Hydrides
- (ii) What is the value of  $\Delta S_{\text{surr.}}$  for the following reaction at 298 K –
- $$6\text{CO}_{2(g)} + 6\text{H}_2\text{O}_{(l)} \longrightarrow \text{C}_6\text{H}_{12}\text{O}_{6(s)} + 6\text{O}_{2(g)}$$
- Given that :  $\Delta G^\circ = 2879 \text{ kJ mol}^{-1}$ ,  $\Delta S = -210 \text{ JK}^{-1} \text{ mol}^{-1}$ .
- (iii) Sucrose decomposes in acid solution to give glucose and fructose according to the first order rate law. The half life of the reaction is 3 hours. Calculate fraction of sucrose which will remain after 8 hours.
- (iv) A solution containing 0.73 g of camphor (molar mass 152 g . mol<sup>-1</sup>) in 36.8 g of acetone (boiling point 56.3°C) boils at 56.55°C. A solution of 0.564 g of unknown compound in the same weight of acetone boils at 56.46°C. Calculate the molar mass of the unknown compound.

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**2014 Edit**



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