

- (ix) The standard potential of the cell in the following reaction is \_\_\_\_\_  
 $\text{Cd}_{(\text{s})} + \text{Cu}_{(\text{aq})}^{2+} \longrightarrow \text{Cd}_{(\text{aq})}^{2+} + \text{Cu}_{(\text{s})}$   
 (  $E_{\text{Cu}^{2+}}^{\circ} = -0.403 \text{ V}$ ,  $E_{\text{Cd}^{2+}}^{\circ} = 0.334 \text{ V}$  )  
 (a)  $-0.737 \text{ V}$  (b)  $0.737 \text{ V}$   
 (c)  $-0.069 \text{ V}$  (d)  $0.069 \text{ V}$
- (x) The value of  $[\text{H}_3\text{O}^+]$  in  $\text{mol lit}^{-1}$  of  $0.001 \text{ M}$  acetic acid solution ( $K_a = 1.8 \times 10^{-5}$ ) is \_\_\_\_\_  
 (a)  $1.34 \times 10^{-1}$  (b)  $1.34 \times 10^{-2}$   
 (c)  $1.34 \times 10^{-3}$  (d)  $1.34 \times 10^{-4}$

**Q.2. Answer the following questions:**

**[8]**

- Write the product formed when alkyl halide reacts with silver nitrite.
- Write the name of product formed, when acetone is treated with 2, 4-dinitrophenyl hydrazine.
- Write the name of biodegradable polyamide copolymer.
- Identify the molecularity of following elementary reaction:  
 $\text{NO}_{(\text{g})} + \text{O}_{(\text{g})} \longrightarrow \text{NO}_{2(\text{g})} + \text{O}_{(\text{g})}$
- What is the action of selenium on magnesium metal?
- Write the name of isomerism in the following complexes:  
 $[\text{Cu}(\text{NH}_3)_4]$   $[\text{PtCl}_4]$  and  $[\text{Pt}(\text{NH}_3)_4]$   $[\text{CuCl}_4]$
- Write the name of the alloy used in Fischer Tropsch process in the synthesis of gasoline.
- Henry's law constant for  $\text{CH}_3\text{Br}_{(\text{g})}$  is  $0.159 \text{ mol dm}^{-3} \text{ bar}^{-1}$  at  $25^\circ\text{C}$ . What is solubility of  $\text{CH}_3\text{Br}_{(\text{g})}$  in water at same temperature and partial pressure of  $0.164 \text{ bar}$ ?

#### SECTION – B

**Attempt any EIGHT of the following questions:**

**[16]**

- Explain pseudo-first order reaction with suitable example.
- Write the consequences of Schottky defect with reasons.
- What is the action of following on ethyl bromide:  
 (i) Na in dry ether  
 (ii) Mg in dry ether
- Explain formation of peptide linkage in protein with an example.
- Derive an expression to calculate molar mass of non volatile solute by osmotic pressure measurement.
- Explain monodentate and ambidentate ligands with example.
- Explain the trends in the following atomic properties of group 16 elements:  
 (i) Atomic radii  
 (ii) Ionisation enthalpy  
 (iii) Electronegativity  
 (iv) Electron gain enthalpy
- Write preparation of phenol from aniline.
- Write chemical reactions to prepare ethanamine from:  
 (i) acetonitrile  
 (ii) nitroethane
- Identify A and B from the following reaction:



# Hsc 2013 Chemistry Solved Paper

**Tobias Bleicker**



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