

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HBr~~ **hydrobromic acid**
- ~~HNO₃~~ **nitric acid**
- ~~H₂SO₄~~ **sulfuric acid**
- ~~HClO₄~~ **perchloric acid**
- ~~Sodium hydroxide~~ **NaOH**
- ~~potassium hydroxide~~ **KOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~barium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **NaOH**

~~HClO₃ is a STRONG acid & it does not completely dissociate in water.~~

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

~~HBr~~

~~HNO₃~~

~~Mg(OH)₂~~

NaOH

NaOH

~~NaOH~~

NaOH

3. What is the formula for hydronium? **H₃O⁺**

4. Cu_2CO_3 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{Cu}_2\text{CO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3 + \text{Cu}_2\text{CO}_3$
- Would you expect a large or small equilibrium constant for this reaction? **small, because Cu_2CO_3 is a weak acid.**
- Would you define Cu_2CO_3 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is Cu_2CO_3 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of Cu_2CO_3 : **Cu_2CO_3**
- Identify Cu_2CO_3 as polyprotic, amphoteric, or neither: **polyprotic, because it contains multiple hydrogens.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{F}^- + \text{H}_2\text{O} \rightarrow \text{HF} + \text{OH}^-$
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither: **neither**

6. HSO_4^- will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{HSO}_4^- + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4 + \text{SO}_4^{2-}$
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO_4^- is a weak acid.**
- Would you define HSO_4^- as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HSO_4^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO_4^- : **SO_4^{2-}**
- Identify HSO_4^- as polyprotic, amphoteric, or neither: **amphoteric, because it has a negative charge (and will like to have a proton) and it has hydrogens (and can lose an acid).**

19 3 Acids And Bases Answer Key Chemistry

Minjie Lin



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Prentice Hall Chemistry, 2000 **General Chemistry** Ralph H. Petrucci, William S. Harwood, 1993 General Chemistry Principles and Modern Applications is recognized for its superior problems lucid writing and precision of argument This updated and expanded edition retains the popular and innovative features of previous editions including Feature Problems follow up Integrative and Practice Exercises to accompany every in chapter Example and Focus On application boxes as well as new Keep in Mind marginal notes Topics covered include atoms and the atomic theory chemical compounds and reactions gases Thermochemistry electrons in atoms chemical bonding liquids solids and intermolecular forces chemical kinetics principles of chemical equilibrium acids and bases electrochemistry representative and transitional elements and nuclear and organic chemistry For individuals interested in a broad overview of chemical principles and applications Cehmistry Textbook for College and University USA Ibrahim Sikder, 2023-06-04 Cehmistry Textbook USA **A Dictionary of Chemistry** Henry Watts, 1865 Physical Chemistry for Engineering and Applied Sciences Frank R. Foulkes, 2012-09-12 Physical Chemistry for Engineering and Applied Sciences is the product of over 30 years of teaching first year Physical Chemistry as part of the Faculty of Applied Science and Engineering at the University of Toronto Designed to be as rigorous as compatible with a first year student s ability to understand the text presents detailed step by step derivations of the equations that permit the student to follow the underlying logic and of equal importance to appreciate any simplifying assumptions made or mathematical tricks employed In addition to the 600 exercises and end of chapter problems the text is rich in worked non trivial examples many of which are designed to be inspiring and thought provoking Step by step derivation of all equations enables the student to smoothly follow the derivation by sight and can be understood relatively easily by students with moderate skills and backgrounds in mathematics Clear and accessible Physical Chemistry for Engineering and Applied Sciences includes The answers to all of the 112 worked examples 99 exercises following many of the worked examples and 496 end of chapter problems Topics not normally seen in introductory physical chemistry textbooks ionic reaction rates activities and activity coefficients or not regularly explained in much detail electrochemistry chemical kinetics with an eye on industrial applications Special appendices that provide detailed explanations of basic integration and natural logarithms for students lacking a background in integral calculus An in depth chapter on electrochemistry in which activities and activity coefficients are used extensively as required for accurate calculations **Watts' Dictionary of Chemistry** Henry Watts, Henry Forster Morley, Matthew Moncrieff Pattison Muir, 1894 Living Science Chemistry 10 Arun Syamal, Living Science for Classes 9 and 10 have been prepared on the basis of the syllabus developed by the NCERT and adopted by the CBSE and many other State Education Boards Best of both the traditional courses and the recent innovations in the field of basic Chemistry have been incorporated The books contain a large number of worked out examples illustrations illustrative questions numerical problems figures tables and graphs **A Dictionary of Chemistry and the**

Allied Branches of Other Sciences Henry Watts, 1883 **Organic Chemistry** David R. Klein, 2022 Organic Chemistry 4th Edition provides a comprehensive yet accessible treatment of all the essential organic chemistry concepts covered in a two semester course Presented with a skills based approach that bridges the gap between organic chemistry theory and real world practice the book places special emphasis on developing their problem solving skills through applied exercises and activities It incorporates Klein s acclaimed SkillBuilder program which contains a solved problem that demonstrates a skill and several practice problems of varying difficulty levels including conceptual and cumulative problems that challenge students to apply the skill in a slightly different environment An up to date collection of literature based problems exposes students to the dynamic and evolving nature of organic chemistry and its active role in addressing global challenges The text is also enriched with numerous hands on activities and real world examples that help students understand both the why and the how behind organic chemistry

Cumulated Index Medicus ,1995 Electrochemistry in Nonaqueous Solutions Kosuke Izutsu, 2009-09-22 An excellent resource for all graduate students and researchers using electrochemical techniques After introducing the reader to the fundamentals the book focuses on the latest developments in the techniques and applications in this field This second edition contains new material on environmentally friendly solvents such as room temperature ionic liquids

High Temperature Polymer Electrolyte Membrane Fuel Cells Qingfeng Li, David Aili, Hans Aage Hjuler, Jens Oluf Jensen, 2015-10-15 This book is a comprehensive review of high temperature polymer electrolyte membrane fuel cells PEMFCs PEMFCs are the preferred fuel cells for a variety of applications such as automobiles cogeneration of heat and power units emergency power and portable electronics The first 5 chapters of the book describe rationalization and illustration of approaches to high temperature PEM systems Chapters 6-13 are devoted to fabrication optimization and characterization of phosphoric acid doped polybenzimidazole membranes the very first electrolyte system that has demonstrated the concept of and motivated extensive research activity in the field The last 11 chapters summarize the state of the art of technological development of high temperature PEMFCs based on acid doped PBI membranes including catalysts electrodes MEAs bipolar plates modelling stacking diagnostics and applications

The American chemist ,1874

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