

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-----------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) for the reaction(s) described. Show the calculation for pH or percent dissociation when required. Write the chemical equation.

1. What substance is H^+ in water? Hydrogen ion is H^+ in water.
2. What substance is OH^- in water? Hydroxide ion is OH^- in water.
3. What substance is H_2O in water? Water is H_2O in water.
4. What substance is H_2O in water? Water is H_2O in water.
5. What substance is H_2O in water? Water is H_2O in water.
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19. What substance is H_2O in water? Water is H_2O in water.
20. What substance is H_2O in water? Water is H_2O in water.

1. H^+
2. OH^-
3. H_2O
4. H_2O
5. H_2O
6. H_2O
7. H_2O
8. H_2O
9. H_2O
10. H_2O
11. H_2O
12. H_2O
13. H_2O
14. H_2O
15. H_2O
16. H_2O
17. H_2O
18. H_2O
19. H_2O
20. H_2O

19 2 Hydrogen Ions And Acidity Work Sheet Answers

Rosina Ehmann



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