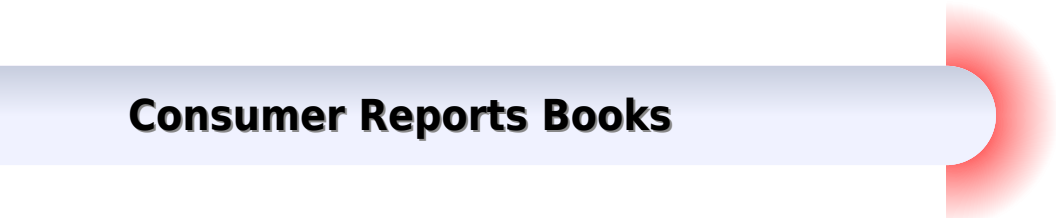




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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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 $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$ 26.982 g $\text{Al} \times 4 \text{ mol Al} \times 1 \text{ mol Al}_2\text{O}_3$ c determine the percent yield
yield actual $\times 100 = 260.0 \text{ g} \times 100 = 86.01\%$ theoretical 302.3 g

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weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12
c is the limiting reagent 4.33 g of h 2 are left over 13

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