

* Show that, the entropy remains constant during a reversible cyclic change.

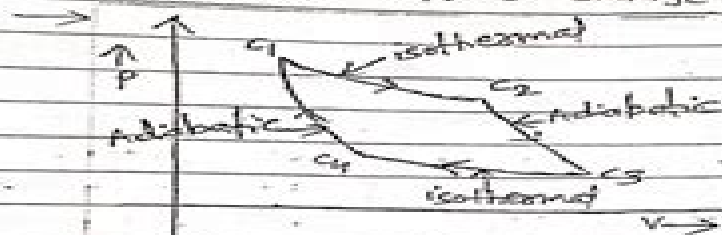


Fig: Carnot's cycle

Consider a reversible cycle $C_1 C_2 C_3 C_4 C_1$ as shown in fig. It consists of 4 operations
 1) From C_1 to C_2 heat Q_1 is absorbed by the system at temp. T_1 . The change in entropy of the system is S_1 . The change is positive.

2) From C_2 to C_3 , no heat is absorbed or rejected by the system, hence no change in entropy

3) From C_3 to C_4 , heat Q_2 is rejected by the system at temp. T_2 . The change in entropy is S_2 . The change is negative.

4) From C_4 to C_1 , no heat is absorbed or rejected by the system, hence no change in entropy.

$\therefore$ Total change in entropy of a working substance during reversible cycle is $\frac{Q_1}{T_1} = \frac{Q_2}{T_2} \Rightarrow \frac{Q_1}{T_1} - \frac{Q_2}{T_2} = 0$

change in entropy of the system of source & sink is zero. Thus during reversible cycle, the entropy of the universe remains constant. i.e. $\Delta S_U = 0$.

* state & explain the principle of increase of entropy or explain entropy of universe in irreversible processes.

Efficiency of Carnot's cycle is

$$\eta = 1 - \frac{Q_2}{Q_1}$$

But, this efficiency is less than that of a reversible engine working between the same two temp's T_1 & T_2 . for which, efficiency

$$\eta = 1 - \frac{T_2}{T_1}$$

$$\therefore 1 - \frac{Q_2}{Q_1} < 1 - \frac{T_2}{T_1}$$

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Table of Contents Applied Thermodynamics Notes

1. Understanding the eBook Applied Thermodynamics Notes
 - The Rise of Digital Reading Applied Thermodynamics Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Applied Thermodynamics Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Applied Thermodynamics Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Applied Thermodynamics Notes
 - Personalized Recommendations

- Applied Thermodynamics Notes User Reviews and Ratings
- Applied Thermodynamics Notes and Bestseller Lists
- 5. Accessing Applied Thermodynamics Notes Free and Paid eBooks
 - Applied Thermodynamics Notes Public Domain eBooks
 - Applied Thermodynamics Notes eBook Subscription Services
 - Applied Thermodynamics Notes Budget-Friendly Options
- 6. Navigating Applied Thermodynamics Notes eBook Formats
 - ePub, PDF, MOBI, and More
 - Applied Thermodynamics Notes Compatibility with Devices
 - Applied Thermodynamics Notes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Applied Thermodynamics Notes
 - Highlighting and Note-Taking Applied Thermodynamics Notes
 - Interactive Elements Applied Thermodynamics Notes
- 8. Staying Engaged with Applied Thermodynamics Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Applied Thermodynamics Notes
- 9. Balancing eBooks and Physical Books Applied Thermodynamics Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Applied Thermodynamics Notes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Applied Thermodynamics Notes
 - Setting Reading Goals Applied Thermodynamics Notes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Applied Thermodynamics Notes
 - Fact-Checking eBook Content of Applied Thermodynamics Notes

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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