# **Energy and Chemical Change**

#### Section 15.1 Energy

pages 516-522

#### **Practice Problems**

pages 519-521

 A fruit and oatmeal bar contains 142 nutritional Calories. Convert this energy to calories.

142 Calories = 142 kcal

142 keaf × 1000 cal 1 keaf = 142,000 cal

An exothermic reaction releases 86.5 kJ. How many kilocalories of energy are released?

 $86.5 \text{ Jef} \times \frac{1 \text{ kcal}}{4.184 \text{ kef}} = 20.7 \text{ kcal}$ 

3. Challenge Define a new energy unit, named after yourself, with a magnitude of one-tenth of a calorie. What conversion factors relate this new unit to joules? To Calories?

Unit X = 0.1 cal 1 cal = 4.184 J

X = (0.1 cal)(4.184 J/cal) = 0.4184 J

1 cal = 0.001 Calorie

X = {0.1 cal}(1 Cal/1000 cal) = 0.0001 Calorie

 If the temperature of 34.4 g of ethanol increases from 25.0°C to 78.8°C, how much heat has been absorbed by the ethanol? Refer to Table 15.2.

 $q = c \times m \times \Delta T$  $q = 2.44 \text{ I/(gr.^{\circ}C)} \times 34.4 \text{ gr} \times 53.8\% = 4.52 \times 10^{3} \text{ J}$ 

5. A 155-g sample of an unknown substance was heated from 25.0°C to 40.0°C. In the process, the substance absorbed 5696 J of energy. What is the specific heat of the substance? Identify the substance among those listed in Table 15.2 on page 520.

 $q = \epsilon \times m \times \Delta T$ 

$$c = \frac{q}{m\Delta T} = \frac{(5696 \text{ J})}{(155 \text{ g})(40.0 - 25.0^{\circ}\text{C})} = 2.45 \text{ J/(g-°C)}$$

The specific heat is very close to the value for ethanol.

6. Challenge A 4.50-g nugget of pure gold absorbed 276 J of heat. The initial temperature was 25.0°C. What was the final temperature?

 $q = c \times m \times \Delta T$ 

$$\Delta T = \frac{q}{cm} = \frac{(276 \text{ J})}{(0.129 \text{ J/g}^{\circ}\text{C})(4.50 \text{ g})} = 475^{\circ}\text{C}$$

$$\Delta T = T_f - T_i = T_f - 25.0^{\circ}C = 475^{\circ}C$$

$$T_{\nu} = 5.00 \times 10^{24} \text{C}$$

#### Section 15.1 Assessment

page 522

 Explain how energy changes from one form to another in an exothermic reaction. In an endothermic reaction.

Chemical potential energy changes to heat in exothermic reactions and the heat is released. In endothermic reactions, heat is absorbed and changed to chemical potential energy.

8. Distinguish between kinetic and potential energy in the following examples: two separated magnets; an avalanche of snow; books on library shelves; a mountain stream; a stock-car race; separation of charge in a battery.

Two separated magnets illustrate potential energy. In a snow avalanche, positional potential energy is changing to kinetic energy. Books on a shelf illustrate positional potential energy. As water races down a mountain stream, positional potential energy is changing to kinetic energy. In a stock-car race, chemical potential energy is being changed to kinetic energy. The separation of charge in a battery illustrates electrical potential energy.

 Explain how the light and heat of a burning candle are related to chemical potential energy.

Chemical potential energy, contained in the candle, is changed to energy in the form of light and heat and released as the chemical combustion reaction takes place.

# **Assessment Energy And Chemical Change Solutions Manual**

United States. Energy Research and Development Administration

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<u>Study Guide/Selected Solutions Manual Julie R. Frentrup, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book</u>

General Organic and Biological Chemistry Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features

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