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Roman Wölfel



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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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1 $\text{K}_3\text{PO}_4 \text{aq}$ + 1 $\text{Al}^{3+} \text{aq}$ + 3 $\text{OH}^- \text{aq}$ + 3 $\text{H}_2\text{O} \text{l}$ ionic $3\text{Cl}^- \text{aq}$ + 1 $\text{Ag}^+ \text{aq}$ + 3 $\text{NH}_4^+ \text{aq}$ + 3 $\text{K}^+ \text{aq}$ net ionic $\text{AgCl} \text{s}$ + 2 $\text{H}_2\text{O} \text{l}$ + 3 $\text{mg} \text{s}$

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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