

19.5

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid-base indicator |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the requested unit or numerical information of the values provided that are indicated in parentheses. Significant figures are indicated with the appropriate symbol.

1. What substance is _____, a weak hydrogen ion donor?
2. What substance is _____, the conjugate base of that which provides an effective buffer system _____?
3. What is the _____ of a weak acid _____?
4. What is the _____ of a weak acid _____?
5. What is the _____ of a weak acid _____?
6. What is the _____ of a weak acid _____?
7. What is the _____ of a weak acid _____?
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16. What is the _____ of a weak acid _____?
17. What is the _____ of a weak acid _____?
18. What is the _____ of a weak acid _____?
19. What is the _____ of a weak acid _____?
20. What is the _____ of a weak acid _____?

1. _____
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20. _____

19 2 Hydrogen Ions And Acidity Work Sheet Answers

Hussin A.Rothana



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