

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HBr~~ **hydrobromic acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO₃~~ **nitric acid**
- ~~HClO₄~~ **perchloric acid**
- ~~Sulfuric hydrosulfate~~ **H₂SO₄**
- ~~potassium hydroxide~~ **KOH**
- ~~lithium hydroxide~~ **LiOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **NH₄OH**

HClO₃ is NOT a strong acid & does not completely dissociate in water.

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

~~HBr~~

~~HNO₃~~

~~Mg(OH)₂~~

HNO₂

NaOH

~~NaOH~~

H₂SO₄

3. What is the formula for hydronium? **H₃O⁺**

4. Cu_2CO_3 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{Cu}_2\text{CO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3 + \text{Cu}_2\text{CO}_3$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because Cu_2CO_3 is a weak acid.**
- Would you define Cu_2CO_3 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is Cu_2CO_3 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of Cu_2CO_3 : **Cu_2CO_3**
- Identify Cu_2CO_3 as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple hydrogens.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{F}^- + \text{H}_2\text{O} \rightarrow \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither. **neither**

6. HSO_4^- will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{HSO}_4^- + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4 + \text{SO}_4^{2-}$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO_4^- is a weak acid.**
- Would you define HSO_4^- as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HSO_4^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO_4^- : **SO_4^{2-}**
- Identify HSO_4^- as polyprotic, amphoteric, or neither. **amphoteric, because it has a negative charge (and will like to have a proton) and it has hydrogens (and can lose an H^+).**

19 3 Acids And Bases Answer Key Chemistry

David R. Klein



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Sciences Henry Watts,1883 **Organic Chemistry** David R. Klein,2022 Organic Chemistry 4th Edition provides a comprehensive yet accessible treatment of all the essential organic chemistry concepts covered in a two semester course Presented with a skills based approach that bridges the gap between organic chemistry theory and real world practice the book places special emphasis on developing their problem solving skills through applied exercises and activities It incorporates Klein s acclaimed SkillBuilder program which contains a solved problem that demonstrates a skill and several practice problems of varying difficulty levels including conceptual and cumulative problems that challenge students to apply the skill in a slightly different environment An up to date collection of literature based problems exposes students to the dynamic and evolving nature of organic chemistry and its active role in addressing global challenges The text is also enriched with numerous hands on activities and real world examples that help students understand both the why and the how behind organic chemistry

Cumulated Index Medicus ,1971 Electrochemistry in Nonaqueous Solutions Kosuke Izutsu,2009-09-22 An excellent resource for all graduate students and researchers using electrochemical techniques After introducing the reader to the fundamentals the book focuses on the latest developments in the techniques and applications in this field This second edition contains new material on environmentally friendly solvents such as room temperature ionic liquids

High Temperature Polymer Electrolyte Membrane Fuel Cells Qingfeng Li,David Aili,Hans Aage Hjuler,Jens Oluf Jensen,2015-10-15 This book is a comprehensive review of high temperature polymer electrolyte membrane fuel cells PEMFCs PEMFCs are the preferred fuel cells for a variety of applications such as automobiles cogeneration of heat and power units emergency power and portable electronics The first 5 chapters of the book describe rationalization and illustration of approaches to high temperature PEM systems Chapters 6-13 are devoted to fabrication optimization and characterization of phosphoric acid doped polybenzimidazole membranes the very first electrolyte system that has demonstrated the concept of and motivated extensive research activity in the field The last 11 chapters summarize the state of the art of technological development of high temperature PEMFCs based on acid doped PBI membranes including catalysts electrodes MEAs bipolar plates modelling stacking diagnostics and applications

The American chemist ,1874

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