

$$\begin{aligned}
 M &= RT \left( \frac{\rho}{p} \right)_{p=0} = (0.08206 \text{ L atm mol}^{-1} \text{ K}^{-1}) \times (273.15 \text{ K}) \times (1.42755 \text{ g L}^{-1} \text{ atm}^{-1}) \\
 &= \boxed{31.9987 \text{ g mol}^{-1}}
 \end{aligned}$$

The value obtained for  $R$  deviates from the accepted value by 0.005 per cent. The error results from the fact that only three data points are available and that a linear extrapolation was employed. The molar mass, however, agrees exactly with the accepted value, probably because of compensating plotting errors.

**E1.11(b)** The mass density  $\rho$  is related to the molar volume  $V_m$  by

$$V_m = \frac{M}{\rho}$$

where  $M$  is the molar mass. Putting this relation into the perfect gas law yields

$$pV_m = RT \quad \text{so} \quad \frac{pM}{\rho} = RT$$

Rearranging this result gives an expression for  $M$ ; once we know the molar mass, we can divide by the molar mass of phosphorus atoms to determine the number of atoms per gas molecule

$$M = \frac{RT\rho}{p} = \frac{(62.364 \text{ L Torr K}^{-1} \text{ mol}^{-1}) \times [(100 + 273) \text{ K}] \times (0.6388 \text{ g L}^{-1})}{120 \text{ Torr}} = 124 \text{ g mol}^{-1}$$

The number of atoms per molecule is

$$\frac{124 \text{ g mol}^{-1}}{31.0 \text{ g mol}^{-1}} = 4.00$$

suggesting a formula of  $\boxed{\text{P}_4}$

**E1.12(b)** Use the perfect gas equation to compute the amount; then convert to mass.

$$pV = nRT \quad \text{so} \quad n = \frac{pV}{RT}$$

We need the partial pressure of water, which is 53 per cent of the equilibrium vapour pressure at the given temperature and standard pressure.

$$\begin{aligned}
 p &= (0.53) \times (2.69 \times 10^3 \text{ Pa}) = 1.43 \times 10^3 \text{ Pa} \\
 \text{so } n &= \frac{(1.43 \times 10^3 \text{ Pa}) \times (250 \text{ m}^3)}{(8.3145 \text{ J K}^{-1} \text{ mol}^{-1}) \times (23 + 273) \text{ K}} = 1.43 \times 10^2 \text{ mol}
 \end{aligned}$$

$$\text{or } m = (1.43 \times 10^2 \text{ mol}) \times (18.0 \text{ g mol}^{-1}) = 2.61 \times 10^3 \text{ g} = \boxed{2.61 \text{ kg}}$$

**E1.13(b)** (a) The volume occupied by each gas is the same, since each completely fills the container. Thus solving for  $V$  from eqn 1.4 we have (assuming a perfect gas)

$$\begin{aligned}
 V &= \frac{n_J RT}{p_J} \quad n_{\text{Ne}} = \frac{0.225 \text{ g}}{20.18 \text{ g mol}^{-1}} \\
 &= 1.113 \times 10^{-2} \text{ mol}, \quad p_{\text{Ne}} = 66.5 \text{ Torr}, \quad T = 300 \text{ K}
 \end{aligned}$$

$$V = \frac{(1.113 \times 10^{-2} \text{ mol}) \times (62.36 \text{ L Torr K}^{-1} \text{ mol}^{-1}) \times (300 \text{ K})}{66.5 \text{ Torr}} = 3.137 \text{ L} = \boxed{3.14 \text{ L}}$$

# Physical Chemistry Atkins 7th Edition Solutions Manual

**Brendan G. Carr**



## **Physical Chemistry Atkins 7th Edition Solutions Manual:**

*Student's Solutions Manual to Accompany Atkins' Physical Chemistry* Peter William Atkins, 2002-01 This solutions manual provides the authors detailed solutions to exercises and problems in the seventh edition of Physical Chemistry by Peter Atkins and Julio de Paula The manual is intended for students and instructors alike and comprises solutions to the A exercises at the end of each chapter solutions to selected numerical theoretical and additional problems at the end of each chapter helpful comments that aid the student s understanding of selected solutions friendly guidance from the authors in the working of each solution *US Solutions Manual to Accompany Elements of Physical Chemistry 7e* David (Faculty Education Director and Undergraduate Dean for the Faculty of Science Smith, and Deputy Head of the School of Chemistry Bristol University), David Smith, 2017-10-05 The Solutions Manual to Accompany Elements of Physical Chemistry 7th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book *Student Solutions Manual to Accompany Atkins' Physical Chemistry* Charles Trapp, M. P. Cady, Carmen Giunta, 2014 The Student Solutions Manual to accompany Atkins Physical Chemistry 10th edition provides full worked solutions to the a exercises and the odd numbered discussion questions and problems presented in the parent book The manual is intended for students and instructors alike and provides helpful comments and friendly advice to aid understanding *US Solutions Manual to Accompany Elements of Physical Chemistry 7e* David Smith, 2017-09-28 The Solutions Manual to Accompany Elements of Physical Chemistry 7th edition contains full worked solutions to all end of chapter discussion questions and exercises featured in the book The manual provides helpful comments and friendly advice to aid understanding It is also a valuable resource for any lecturer who wishes to use the extensive selection of exercises featured in the text to support either formative or summative assessment and wants labour saving ready access to the full solutions to these questions

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