



Activity B: Classifying reactions

Get the Gizmo ready:

- Turn off Show summary and Show histograms.



Introduction: Chemical equations show how **compounds** and **elements** react with one another. An element is a substance consisting of one kind of atom, such as aluminum (Al) or oxygen gas (O_2). A compound is a substance made of more than one kind of atom, such as water (H_2O) or table salt (NaCl).

Question: How are chemical reactions classified?

1. **Match:** Most chemical reactions can be classified as one of four types. Using the chemical equations in the Gizmo as a guide, match the following definitions to the type of reaction.

B	One reactant is broken down into two or more products.
E	A fuel combines with oxygen to produce carbon dioxide (CO_2) and water (H_2O).
A	Two or more reactants combine to form one product.
D	Two compounds react to form two different compounds.
C	A compound reacts with an element to form a new compound and a different element.

A. Synthesis
B. Decomposition
C. Single replacement
D. Double replacement
E. Combustion

2. **Practice:** Balance each of the chemical equations below. (Some equations may already be in balance.) In the space to the right, classify the reaction as a *synthesis*, *decomposition*, *single replacement*, *double replacement*, or *combustion*.

A.	1	$\text{AgNO}_3 +$	1	$\text{KCl} \rightarrow$	1	$\text{AgCl} +$	1	KNO_3	Two replacements
B.	1	$\text{H}_2\text{O} +$	1	$\text{SO}_3 \rightarrow$	1	H_2SO_4			Synthesis
C.	2	$\text{KI} +$	1	$\text{Cl}_2 \rightarrow$	2	$\text{KCl} +$	1	I_2	Single replacement
D.	2	$\text{NaHCO}_3 \rightarrow$	1	$\text{Na}_2\text{CO}_3 +$	1	$\text{H}_2\text{O} +$	1	CO_2	Decomposition
E.	1	$\text{Zn} +$	2	$\text{HCl} \rightarrow$	1	$\text{ZnCl}_2 +$	1	H_2	Single replacement

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C Cleary

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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