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## FOCUS 2

### MOLECULES

**2A.2** (a) 5; (b) 8; (c) 12; (d) 7

**2A.4** (a)  $[\text{Ar}] 3d^1$ ; (b)  $[\text{Ar}] 3d^5$ ; (c)  $[\text{Kr}] 4d^{10}$ ; (d)  $[\text{Xe}]$

**2A.6** (a)  $[\text{Kr}]$ ; (b)  $[\text{Xe}] 4f^{14}5d^5$ ; (c)  $[\text{Xe}]$ ; (d)  $[\text{Ar}]$

**2A.8** (a) Ca:  $[\text{Ar}] 4s^2$ ,  $\text{Ti}^{2+}$ :  $[\text{Ar}] 3d^2$ ;  $\text{V}^{3+}$ :  $[\text{Ar}] 3d^2$ . In the d block, the energies of the  $n-1$  d-orbitals lie below those of the ns-orbitals. Therefore, when V and Ti form ions they lose their 3s electrons before losing their 4d electrons. (b) Ca: no unpaired electrons;  $\text{Ti}^{2+}$ : two unpaired electrons;  $\text{V}^{3+}$ : two unpaired electrons. (c)  $\text{Ti}^{2+}$ :  $[\text{Ar}] 3d^1$  no neutral atom has this electron configuration.

**2A.10** (a)  $\text{Au}^{3+}$ ; (b)  $\text{Os}^{3+}$ ; (c)  $\text{I}^{3+}$ ; (d)  $\text{As}^{3+}$

**2A.12** (a)  $\text{Cr}^{2+}$ ; (b)  $\text{Ag}^{2+}$ ; (c)  $\text{Zn}^{2+}$ ; (d)  $\text{Pb}^{2+}$

**2A.14** (a) 4d; (b) 3p; (c) 5p; (d) 4s

**2A.16** (a) +1; (b) -2; (c) +2; (d) -3; (e) -1

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**N Colangelo**



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