

Thermochemistry Review

Multiple Choice

Identify the choice that best completes the statement or answers the question.

Use the following information to answer the next questions.

Decaying matter may provide energy directly through combustion or indirectly through decomposition to form methane, which can then be burned.



- _____ 1. The burning of methane is an example of an
- exothermic reaction in which carbon is oxidized.
 - exothermic reaction in which carbon is reduced.
 - endothermic reaction in which carbon is oxidized.
 - endothermic reaction in which carbon is reduced.
- _____ 2. The molar enthalpy of combustion for methane is
- 560.5 kJ/mol
 - 710.1 kJ/mol
 - 802.3 kJ/mol
 - 951.9 kJ/mol
- _____ 3. An equation that represents an endothermic change is
- $\text{C}_2\text{H}_5\text{OH}(\text{l}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$
 - $2 \text{C}(\text{s}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g})$
 - $\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$
 - $2 \text{C}(\text{s}) + 3 \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$
- _____ 4. Sulfuric acid is used to dissolve copper ore in many mining operations. Sulfuric acid may be produced on site by burning sulfur and dissolving the products in water. The overall equation for this reaction is



In this process, what is the enthalpy of reaction for the production of one mole of sulfuric acid?

- +4225.6 kJ
 - +528.2 kJ
 - 528.2 kJ
 - 4225.6 kJ
- _____ 5. Which of the following statements is **false**?
- Light energy is converted into chemical energy during photosynthesis.
 - Chemical energy is converted into heat and light energy during the combustion of hydrocarbons.
 - Light energy is converted into chemical energy to produce a firefly's glow.
 - Chemical energy is converted into electrical energy as a voltaic cell discharges.

Multiple Choice Questions On Thermochemistry

RC Schank



Multiple Choice Questions On Thermochemistry:

Thermochemistry Mr. Ajeet Singh, 2018-05-16 Introduces thermodynamic principles related to chemical processes including enthalpy entropy Gibbs free energy and calorimetry Internal Combustion Engines Shyam K. Agrawal, 2006 Salient Features The New Edition Is A Thoroughly Revised Version Of The Earlier Edition And Presents A Detailed Exposition Of The Basic Principles Of Design Operation And Characteristics Of Reciprocating I C Engines And Gas Turbines Chemistry Of Combustion Engine Cooling And Lubrication Requirements Liquid And Gaseous Fuels For Ic Engines Compressors Supercharging And Exhaust Emission Its Standards And Control Thoroughly Explained Jet And Rocket Propulsion Alternate Potential Engines Including Hybrid Electric And Fuel Cell Vehicles Are Discussed In Detail Chapter On Ignition System Includes Electronic Injection Systems For Si And Ci Engines 150 Worked Out Examples Illustrate The Basic Concepts And Self Explanatory Diagrams Are Provided Throughout The Text More Than 200 Multiple Choice Questions With Answers A Good Number Of Review Questions Numerical With Answers For Practice Will Help Users In Preparing For Different Competitive Examinations With These Features The Present Text Is Going To Be An Invaluable One For Undergraduate Mechanical Engineering Students And Amie Candidates **Engineering Chemistry I (for BPUT)** Patrizia Bocchetta, 2024

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