

CHEM 101 Worksheet, Chapter 3: Mass Relationships in Chemical Reactions

Multiple choice questions: circle the correct answer(A-E)

Show your work questions (p. 2): show all work! Use the amu for each element from the periodic table!

$N_A = 6.02 \times 10^{23}$

1. Atom X has two isotopes: 25.0 wt% is ^{37}X and 75.0% is ^{36}X . What is the average molar mass of X?

- A. 36.0 amu B. 72.0 amu C. 36.5 amu D. 35.5 amu E. 34.5 amu

2. The molar mass (in amu) of $\text{ZnSO}_4 \cdot 6\text{H}_2\text{O}$ =

- A. 161.5 B. 179.5 C. 269.5 D. 53.5 E. 215.5

3. How many atoms of Na in 5.10 moles of Na_2CO_3 ?

- A. 3.07×10^{23} B. 6.14×10^{24} C. 6.14×10^{23} D. 1.84×10^{25} E. 10.2

4. What is the mass in grams of 1.00×10^{12} atoms of lead, Pb?

- A. $3.44 \times 10^{-10}\text{g}$ B. $2.07 \times 10^{114}\text{g}$ C. $1.25 \times 10^{38}\text{g}$ D. $3.44 \times 10^{-34}\text{g}$ E. $1.25 \times 10^{14}\text{g}$

5. 0.2 mol Na_2CO_3 = 4.24 g Na_2CO_3 = 0.051 mol

- A. 530g, 449 mol B. 21.2g, 449 mol C. 530g, 0.051 mol D. 21.2g, 0.040 mol E. 16.6g, 0.051 mol

6. The wt% composition in tin(IV)oxide is 88% Sn 12% O (Sn = tin, element # 50)

- A. 21%Sn, 79%O B. 88%Sn, 12%O C. 12%Sn, 88%O D. 94%Sn, 6%O E. 79% Sn, 21%O

7. A compound has the following wt% composition: 2.1% H, 65.3% O, 32.6% S. What is its empirical formula?

- A. H_2SO_4 B. $\text{H}_4\text{S}_2\text{O}_8$ C. S_4HO_2 D. H_2SO_2 E. HO_4S

8. How many gram Na in 5.00g Na_2SO_4 ?

- A. 0.035 g B. 1.31 g C. 1.62 g D. 0.97 g E. 0.81 g

9. A compound containing only carbon and hydrogen has 85.7 wt% C and 14.3 wt% H. Its molar mass is 56 amu. What is its molecular formula?

- A. C_4H_8 B. C_2H_4 C. C_3H D. CH_2 E. CH

10. Balance the following reactions. No fractional coefficients, coefficients must be the lowest whole numbers.

a) 2 K + 2 H_2O $\rightarrow$ 2 KOH + 1 H_2

b) 2 H_2O_2 $\rightarrow$ 2 H_2O + 1 O_2

c) 2 C_2H_6 + 7 O_2 $\rightarrow$ 4 CO_2 + 6 H_2O

d) 3 KOH + 1 H_3PO_4 $\rightarrow$ 1 K_3PO_4 + 3 H_2O

e) 2 $\text{C}_2\text{H}_5\text{OH}$ + 3 O_2 $\rightarrow$ 4 CO_2 + 6 H_2O

11. How many g of CO_2 are formed from the combustion of 2.9 g of butane, C_4H_{10} , with excess oxygen?

The balanced reaction is: $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$

- A. 0.26g B. 6.6g C. 8.8g D. 10.5g E. 2.2g

12. How many g I_2 will react with 20.4 g Al to form AlI_3 (aluminum iodide)? As always, first balance the reaction! (Hint: this is not a limiting reagent problem)

- A. 144. g B. 288. G C. 192. G D. 72 g E. 576 g

13. For the balanced reaction: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$. How many moles of NH_3 will be produced when 0.50 mol N_2 reacts with 1.0 mol H_2 ?

- A. 0.67 mol B. 1.0 mol C. 0.50 mol D. 2.0 mol E. 3.0 mol

14. Calculate the mass of CO_2 formed when 2.3 g of ethanol, $\text{C}_2\text{H}_5\text{OH}$ reacts with 3.2 g oxygen O_2 to form CO_2 and H_2O . (You balanced the reaction in 8.e)! This is a limiting reagent problem!

- A. 4.6g B. 2.2g C. 4.4g D. 6.6g E. 2.93g

Multiple Choice Question For Chemistry 101

M Mosston



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