CHEM 101 Worksheet, Chapter 6: Thermochemistry

Multiple choice questions 1-10 (page 1): circle the correct answer(A-E)

Show your work questions 11-13 (pages 2-3): give the complete answer in the space provided R = 8.314 J/(mol.K) = 0.08206 L.atm(mol.K) 1 cal = 4.184 J 0 °C = 273 K

Calculate the heat effect q (in kJ) when 112 g of iron (Fe(s)) is heated from 20 °C to 300 °C. The specific heat of Fe(s) = 0.48 J/g.K

A. -34.9 kJ kJ

B. +34.9 kJ

 $C_{\rm s} = 15.0 \; {\rm kJ}$

D. -270 J

E. +15.0 kJ

 The dissolution reaction for ammonium chloride: NH₄Cl(s) → NH₄'(aq) + Cl⁻(aq) is endothermic. When ammonium chloride is dissolved in water in an insulated (adiabatic) system, the temperature will:

A. increase

B. decrease

C. stay the same

3. Consider the reaction: $CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$. $\Delta H_{rm} = +178 \text{ kJ}$ What is the enthalpy change when 20.0 kg CaO is formed? A. $+6.36\times10^s \text{ kJ}$ B. $-6.36\times10^s \text{ kJ}$ C. $+6.36\times10^s \text{ J}$ D. 63.6 kJ E. +178 kJ4. Consider the reaction: $2NO + O_2 \rightarrow 2NO_2$ $\Delta H_{rm} = -114.6 \text{ kJ}$ As always, the ΔH is "for the reaction as written"! Calculate the enthalpy change ΔH , when 12.6 g NO_2 is produced. A. -31.4 kJ B. +31.4 kJ C. -15.7 kJ D. +15.7 kJ E. none of these

5. How much heat is associated with the reaction that produces 256.1 g of ammonia, NH₃? $N_2(g) + 3 H_2(g) \rightarrow 2 NH_2(g) = \Delta H^2_{cm} = -92.45 \text{ kJ}$

A. 92.45 kJ of heat is released B. 696.1 kJ of heat is released C. 184.9 kJ of heat is absorbed D. 184.9 kJ of heat is released E. 92 kJ of heat is released

6. Which two of the following are not formation reactions?

1. $\frac{1}{2}N_{2}(g) + \frac{1}{2}H_{2}(g) \rightarrow NH_{3}(g)$ 2. $C(s) + O_{2}(g) \rightarrow CO_{2}(g)$ 3. $2SO_{2}(g) + O_{2}(g) \rightarrow 2SO_{3}(g)$ 4. $2Fe(s) + \frac{1}{2}O_{2}(g) \rightarrow Fe_{2}O_{3}(s)$ 5. $\frac{1}{2}H_{2}(g) + \frac{1}{2}I_{2}(g) \rightarrow HI(g)$ A. 1 and 2 B. 3 and 4 C. 2 and 3 D. 4 and 5 E. 3 and 5 7. Consider the thermochemical reaction equation: $2CO(g) + O_{2}(g) \rightarrow 2CO_{2}(g) \rightarrow H = -198 \text{ kJ}$ What is the enthalpy change for the reaction: $CO_{2}(g) \rightarrow CO(g) + \frac{1}{2}O_{2}(g)$ A. +198 kJ B. -99 kJ C. +99 kJ D. +396 kJ E. -198 kJ

8. Calculate the enthalpy change for the reaction: $2KClO_3 \rightarrow 2\ KCl + O_2(g)$ Standard enthalpy of formation values: $\Delta H_i^*(KClO_3) = -389\ kJ/mol\ \Delta H_i^*(KCl) = -436\ kJ/mol\ A. -94\ kJ$ B. -47 kJ C. +94 kJ D. +47 kJ

E. Cannot be calculated because ΔH_f°(O₂(g)) is not given

9. Given the following two thermochemical equations:

E. -668 kJ

10. Which of the following statements is correct for a combustion reaction in an adiabatic "bomb" calorimeter?

A. the pressure is constant B. the temperature is constant C the volume is constant D. the calorimeter constant's C units are J/g. E. $\Delta T(K) = \Delta T(^{\circ}C) + 273.15$ Show your work for the following questions.

Multiple Choice Thermochemistry Questions

Neil D. Jespersen, Pamela Kerrigan

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