

Multiple Choice: Equilibria Theory (K_{eq}, K_c, K_p)

1. The equilibrium constant $K = \frac{[R]^2 [S]}{[P]^3}$ would apply to:

- A) $2R_{(aq)} + S_{(aq)} \rightleftharpoons 3P_{(aq)}$ B) $3P_{(aq)} \rightleftharpoons 2R_{(aq)} + S_{(aq)}$ C) $2R_{(aq)} + S_{(aq)} \rightleftharpoons 3P_{(aq)} + P_{(aq)}$
D) $3P_{(aq)} \rightleftharpoons 2R_{(aq)} + S_{(aq)}$ E) more than one of the above

ANS: B

1. Choose the FALSE statement regarding the equilibrium state:

- A) rate of the reverse reaction = rate of forward reaction B) can only occur in a open system
C) concentrations of all species remain constant D) can only occur in a closed system

ANS: B

1. The numerical value of the equilibrium constant for a reaction can tell you:

- A) how long it takes to reach equilibrium
B) the extent to which the forward reaction occurs
C) whether a reaction is endothermic or exothermic
D) whether a temperature increase favors products or reactants
E) more than one of the above

ANS: B

46. Which will cause an increase in the numerical value of the equilibrium constant for this reaction?



- A) decreasing the temperature C) finely powdering the C(s)
B) decreasing the volume D) increasing $[H_2]$

ANS: A

48. Which statement is true for this reaction if its equilibrium constant equals 1.6×10^6 ?



- A) It produces mostly products B) It goes only in one direction
C) It does not take place to a significant extent D) It takes place very rapidly

ANS: A

37. This reaction is allowed to come to equilibrium: $2NO(g) + 2CO(g) \rightarrow N_2(g) + 2CO_2(g)$

Additional NO(g) is then injected into the reaction, and equilibrium is restored. Which statement is true?

- A) All concentrations remain the same
B) A new equilibrium is reached such that $[CO]$ decreases
C) A new equilibrium is reached such that $[N_2]$ decreases
D) The numerical value of the equilibrium constant increases

Multiple Choice Concept Questions On Equilibrium

Xiang Xie



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