

Chapter

7

THERMOCHEMISTRY

MULTIPLE CHOICE QUESTIONS

- Spontaneous reactions are: (a) Reversible (b) Irreversible (c) Non irreversible (d) None of these (LHR 05)
- The standard heat changes occur at: (a) 25°C and 2 atm (b) 298 K and 1 atm (c) 25°C and 1 mm hg (d) 273 K and 1 atm (BWP 08, GRW 06)
- $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ the change in enthalpy for reaction is called: (a) Heat of reaction (b) heat of formation (c) Heat of neutralization (d) Heat of combustion (LHR 06)
- The net change in energy in a chemical reaction is same whether it takes place directly or indirectly. It is called: (a) Henry's law (b) Charlie's law (c) Hess's law (d) Graham's law (RWP 11, GRW 07, BWP 10)
- The enthalpy change when one mole of substance is completely burnt in excess of oxygen is called: (a) Enthalpy of atomization (b) Enthalpy of neutralization (c) Enthalpy of Combustion (d) Enthalpy of formation (LHR 07)
- Enthalpy change for the reaction: $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ is called enthalpy of: (a) Formation (b) Combustion (c) Neutralization (d) Atomization (LHR 08)
- Standard enthalpies are measured at: (a) 273 K (b) 298K (c) 373 K (d) All of these (FSD 11, LHR 09)
- The exothermic process is: (a) Evaporation (b) Sublimation (c) Respiration (d) Boiling (DGK 10, GRW, LHR 11)
- A state function which describes together the internal energy and the product of pressure and volume is called: (a) Enthalpy (b) Internal energy (c) Work (d) Free energy (GRW 11)

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Sharma K.K. & Sharma L.K.



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Engineering Chemistry I (for BPUT) Patrizia Bocchetta, 2024 **IIT Chemistry-II**, A Textbook of Physical Chemistry, 6th Edition Sharma K.K. & Sharma L.K., 2016 A Textbook of Physical Chemistry **AP Chemistry Premium, 2024: 6 Practice Tests + Comprehensive Review + Online Practice** Neil D. Jespersen, Pamela Kerrigan, 2023-07-04

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Engineering Chemistry Shikha Agarwal,2019-05-23 Written in lucid language the book offers a detailed treatment of fundamental concepts of chemistry and its engineering applications

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