

## Electron Configuration Practice - Homework - **KEY**

In the space below, write the expanded electron configurations of the following elements:

- |    |           |  |
|----|-----------|--|
| 1) | sodium    | $1s^2 2s^2 2p^6 3s^1$                        |
| 2) | potassium | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$              |
| 3) | chlorine  | $1s^2 2s^2 2p^6 3s^2 3p^5$                   |
| 4) | bromine   | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$ |
| 5) | oxygen    | $1s^2 2s^2 2p^4$                             |

In the space below, write the abbreviated electron configurations of the following elements:

- |     |           |                         |
|-----|-----------|-------------------------|
| 6)  | manganese | $[\text{Ne}] 3s^2$      |
| 7)  | silver    | $[\text{Kr}] 5s^2 4d^9$ |
| 8)  | nitrogen  | $[\text{He}] 2s^2 2p^3$ |
| 9)  | sulfur    | $[\text{Ne}] 3s^2 3p^4$ |
| 10) | argon     | $[\text{Ne}] 3s^2 3p^6$ |

In the space below, write the orbital notation of the following elements:

- |               |                   |                            |                            |                            |                            |                            |                            |                            |                            |                            |                            |                            |                            |                            |                            |                            |                  |
|---------------|-------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|----------------------------|------------------|
| 11) manganese | 12 e <sup>-</sup> | $\uparrow\downarrow$<br>1s | $\uparrow\downarrow$<br>2s | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>3s |                            |                            |                            |                            |                            |                            |                            |                            |                            |                  |
| 12) gallium   | 31 e <sup>-</sup> | $\uparrow\downarrow$<br>1s | $\uparrow\downarrow$<br>2s | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>3s | $\uparrow\downarrow$<br>3p | $\uparrow\downarrow$<br>3p | $\uparrow\downarrow$<br>3p | $\uparrow\downarrow$<br>4s | $\uparrow\downarrow$<br>3d | $\uparrow\downarrow$<br>3d | $\uparrow\downarrow$<br>3d | $\uparrow\downarrow$<br>3d | $\uparrow\downarrow$<br>3d | $\uparrow$<br>4p |
| 13) nitrogen  | 7 e <sup>-</sup>  | $\uparrow\downarrow$<br>1s | $\uparrow\downarrow$<br>2s | $\uparrow$<br>2p           | $\uparrow$<br>2p           | $\uparrow$<br>2p           |                            |                            |                            |                            |                            |                            |                            |                            |                            |                            |                  |
| 14) sulfur    | 16 e <sup>-</sup> | $\uparrow\downarrow$<br>1s | $\uparrow\downarrow$<br>2s | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>3s | $\uparrow\downarrow$<br>3p | $\uparrow$<br>3p           | $\uparrow$<br>3p           |                            |                            |                            |                            |                            |                            |                  |
| 15) argon     | 18 e <sup>-</sup> | $\uparrow\downarrow$<br>1s | $\uparrow\downarrow$<br>2s | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>2p | $\uparrow\downarrow$<br>3s | $\uparrow\downarrow$<br>3p | $\uparrow\downarrow$<br>3p | $\uparrow\downarrow$<br>3p |                            |                            |                            |                            |                            |                            |                  |

Determine what elements are denoted by the following electron configurations:

- |     |   |                    |
|-----|---|--------------------|
| 16) | $1s^2 2s^2 2p^6 3s^2 3p^4$                        | <b>sulfur</b>      |
| 17) | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ | <b>rubidium</b>    |
| 18) | $[\text{Kr}] 5s^2 4d^{10} 5p^3$                   | <b>antimony</b>    |
| 19) | $[\text{Xe}] 6s^2 4f^{14} 5d^6$                   | <b>osmium</b>      |
| 20) | $[\text{Rn}] 7s^2 5f^{11}$                        | <b>einsteinium</b> |

Determine which of the following electron configurations are not valid:

- |     |  |   |
|-----|--|---|
| 21) | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ | <b>not valid (take a look at "4d")</b>                                |
| 22) | $1s^2 2s^2 2p^6 3s^3 3d^5$                   | <b>not valid (3p comes after 3s)</b>                                  |
| 23) | $[\text{Ra}] 7s^2 5f^6$                      | <b>not valid (radium isn't a noble gas)</b>                           |
| 24) | $[\text{Kr}] 5s^2 4d^{10} 5p^5$              | <b>valid</b>  |
| 25) | $[\text{Xe}]$                                | <b>not valid (an element can't be its own electron configuration)</b> |

# Name Electron Arrangements Ws Answers

**CL Gary**



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