## Atoms & The Periodic Table Unit 3 Test Study Guide

#### Lesson 1: The Atom

- 1. Who was the ancient Greek philosopher that came up with the idea that all things are made of small "uncuttable" particles?
- 2. What is an atom?
- 3. What was John Dalton's contribution to the atomic theory?
- 4. What are the parts of an atom?
  - a. Where can each of these particles be found within an atom?
  - b. How do the electrical charges of protons, neutrons, and electrons compare to each other?
  - c. How do the masses of protons, neutrons, and electrons compare to each other?
- 5. What is an atomic number?
  - a. What does this tell about the structure of an atom?
- 6. What is the atomic mass (mass number)?
  - a. What does this tell about the structure of an atom?
    - i. How can the atomic mass help you find the number of neutrons in an atom?

#### Lesson 2: The Periodic Table

- 1. What are elements?
- 2. What is the Periodic Table of Elements?
  - a. How are elements arranged on the Periodic Table?
- 3. Information found on each square of the Periodic Table:
  - a. Atomic Number (See Lesson 1)
  - b. What is a chemical name?
    - i. Examples
  - c. What is a chemical symbol?
    - Examples
  - d. Atomic Mass (See Lesson 1)
- 4. What are the 3 types of elements?
  - a. Where can these elements be found on the Periodic Table?
  - b. What are the characteristics of these types of elements?
- 5. What is a group?
  - a. What do the elements in a group have in common?
- 6. What is a period?
  - a. What do the elements in a period have in common?

#### Lesson 3: Electrons and Chemical Bonding

- 1. What is a chemical bond?
  - a. Why do atoms form bonds?
    - i. How can you figure out what an atom will do to become "happy"?
- 2. What is a group of atoms that are chemically bonded called?
- What are valence electrons?

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Kerry L. Hull, Barbara Janson Cohen

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## **Table of Contents Atoms Periodic Table Study Guide Answer**

- 1. Understanding the eBook Atoms Periodic Table Study Guide Answer
  - The Rise of Digital Reading Atoms Periodic Table Study Guide Answer
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Atoms Periodic Table Study Guide Answer
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Atoms Periodic Table Study Guide Answer
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Atoms Periodic Table Study Guide Answer
  - Personalized Recommendations
  - Atoms Periodic Table Study Guide Answer User Reviews and Ratings
  - Atoms Periodic Table Study Guide Answer and Bestseller Lists
- 5. Accessing Atoms Periodic Table Study Guide Answer Free and Paid eBooks

- Atoms Periodic Table Study Guide Answer Public Domain eBooks
- Atoms Periodic Table Study Guide Answer eBook Subscription Services
- Atoms Periodic Table Study Guide Answer Budget-Friendly Options
- 6. Navigating Atoms Periodic Table Study Guide Answer eBook Formats
  - o ePub, PDF, MOBI, and More
  - Atoms Periodic Table Study Guide Answer Compatibility with Devices
  - Atoms Periodic Table Study Guide Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Atoms Periodic Table Study Guide Answer
  - Highlighting and Note-Taking Atoms Periodic Table Study Guide Answer
  - Interactive Elements Atoms Periodic Table Study Guide Answer
- 8. Staying Engaged with Atoms Periodic Table Study Guide Answer
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Atoms Periodic Table Study Guide Answer
- 9. Balancing eBooks and Physical Books Atoms Periodic Table Study Guide Answer
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Atoms Periodic Table Study Guide Answer
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Atoms Periodic Table Study Guide Answer
  - Setting Reading Goals Atoms Periodic Table Study Guide Answer
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Atoms Periodic Table Study Guide Answer
  - Fact-Checking eBook Content of Atoms Periodic Table Study Guide Answer
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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