

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and vice versa for weak acids and bases.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-----------------|--------------------------------------|-----------------|
| acid | acid dissociation constant (K_a) | acid ionization |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the appropriate formula for each of the following acids, bases, and salts, showing dissociation reactions, and calculate the pH of each solution.

1. Acetic acid, CH_3COOH , $K_a = 1.8 \times 10^{-5}$
2. Hydrofluoric acid, HF , $K_a = 6.6 \times 10^{-4}$
3. Hydrocyanic acid, HCN , $K_a = 4.9 \times 10^{-10}$
4. Hydrobromic acid, HBr , $K_a = 1.0 \times 10^3$
5. Hydrochloric acid, HCl , $K_a = 1.0 \times 10^7$
6. Hydroiodic acid, HI , $K_a = 1.0 \times 10^9$
7. Hydroperchloric acid, $HClO_4$, $K_a = 1.0 \times 10^4$
8. Hydroperbromic acid, $HBrO_4$, $K_a = 1.0 \times 10^4$
9. Hydroperchloric acid, $HClO_4$, $K_a = 1.0 \times 10^4$
10. Hydroperbromic acid, $HBrO_4$, $K_a = 1.0 \times 10^4$
11. Hydroperchloric acid, $HClO_4$, $K_a = 1.0 \times 10^4$
12. Hydroperbromic acid, $HBrO_4$, $K_a = 1.0 \times 10^4$
13. Hydroperchloric acid, $HClO_4$, $K_a = 1.0 \times 10^4$
14. Hydroperbromic acid, $HBrO_4$, $K_a = 1.0 \times 10^4$
15. Hydroperchloric acid, $HClO_4$, $K_a = 1.0 \times 10^4$
16. Hydroperbromic acid, $HBrO_4$, $K_a = 1.0 \times 10^4$
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18. Hydroperbromic acid, $HBrO_4$, $K_a = 1.0 \times 10^4$
19. Hydroperchloric acid, $HClO_4$, $K_a = 1.0 \times 10^4$
20. Hydroperbromic acid, $HBrO_4$, $K_a = 1.0 \times 10^4$

1. Acetic acid, CH_3COOH
2. Hydrofluoric acid, HF
3. Hydrocyanic acid, HCN
4. Hydrobromic acid, HBr
5. Hydrochloric acid, HCl
6. Hydroiodic acid, HI
7. Hydroperchloric acid, $HClO_4$
8. Hydroperbromic acid, $HBrO_4$
9. Hydroperchloric acid, $HClO_4$
10. Hydroperbromic acid, $HBrO_4$
11. Hydroperchloric acid, $HClO_4$
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15. Hydroperchloric acid, $HClO_4$
16. Hydroperbromic acid, $HBrO_4$
17. Hydroperchloric acid, $HClO_4$
18. Hydroperbromic acid, $HBrO_4$
19. Hydroperchloric acid, $HClO_4$
20. Hydroperbromic acid, $HBrO_4$

19 2 Hydrogen Ions And Acidity Section Review Answers

**Linda Swisher, RN, EdD, Kevin T.
Patton, PhD, Gary A. Thibodeau, PhD**

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candium sample was very similar to our neighbors samples if each group were to have used larger samples of candium the respective differences would have been smaller because the larger the sample size the greater the

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