

# Solubility Curve Worksheet

1) Define solubility.

the maximum amount of solute that can dissolve in a solvent at a given temperature

2) Look at the graph below. In general, how does temperature affect solubility?

as temperature increases solubility increases

3) Which compound is LEAST soluble at 10 °C? KClO<sub>3</sub>

4) How many grams of KCl can be dissolved in 100g of water at 80°C? 50 g

5) How many grams of NaCl can be dissolved in 100g of water at 90°C? 40 g

6) At 40°C, how much KNO<sub>3</sub> can be dissolved in 100g of water? 62 g

7) Which compound shows the least amount of change in solubility from 0°C-100°C?

NaCl

8) At 30°C, 90g of NaNO<sub>3</sub> is dissolved in 100g of water. Is this solution saturated or unsaturated?

unsaturated

9) At 60°C, 72g of NH<sub>4</sub>Cl is dissolved in 100g of water. Is this solution saturated or unsaturated?

supersaturated

10) A saturated solution of KClO<sub>3</sub> is formed from one hundred grams of water. If the saturated solution is cooled from 90°C to 50°C, how many grams of precipitate are formed? 50g-20g=30g

11) A saturated solution of NH<sub>4</sub>Cl is formed from one hundred grams of water. If the saturated solution is cooled from 80°C to 40°C, how many grams of precipitate are formed? 65g-45g=20g

12) Which compounds show a decrease in solubility from 0°C-100°C?

NH<sub>3</sub> and Ce<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>

13) Which compound is the most soluble at 10°C?

KI

14) Which compound (besides Ce<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>) is the least soluble at 50°C? KClO<sub>3</sub>

15) For each of the following solutions, explain how much of the solute will dissolve and how much will remain undissolved at the bottom of the test tube?

a) 120 g of KCl in 100 g of water at 80°C

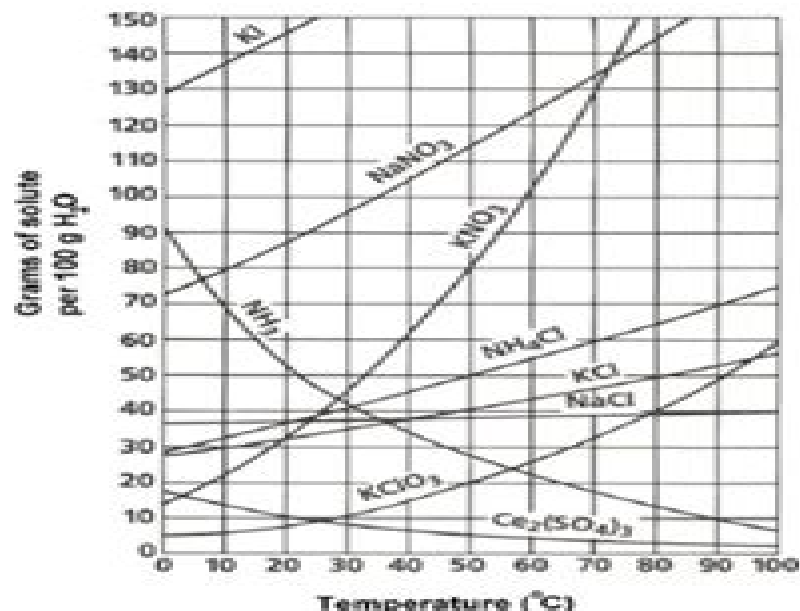
50g will dissolve

120-50=70g remains undissolved

b) 130 g of NaNO<sub>3</sub> in 100 g of water at 50°C

115g will dissolve

130-115=15g undissolved



# Multiple Choice Questions On Solubility Curve

**V. K. Sally**



## Multiple Choice Questions On Solubility Curve:

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Excel With Objective Questions In Chemistry Prof. S. K. Khanna, Dr. N. K. Verma, Dr. B. Kapila, 2006

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### **Table of Contents Multiple Choice Questions On Solubility Curve**

1. Understanding the eBook Multiple Choice Questions On Solubility Curve
  - The Rise of Digital Reading Multiple Choice Questions On Solubility Curve
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions On Solubility Curve
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Questions On Solubility Curve
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions On Solubility Curve
  - Personalized Recommendations
  - Multiple Choice Questions On Solubility Curve User Reviews and Ratings
  - Multiple Choice Questions On Solubility Curve and Bestseller Lists



5. Accessing Multiple Choice Questions On Solubility Curve Free and Paid eBooks
  - Multiple Choice Questions On Solubility Curve Public Domain eBooks
  - Multiple Choice Questions On Solubility Curve eBook Subscription Services
  - Multiple Choice Questions On Solubility Curve Budget-Friendly Options
6. Navigating Multiple Choice Questions On Solubility Curve eBook Formats
  - ePub, PDF, MOBI, and More
  - Multiple Choice Questions On Solubility Curve Compatibility with Devices
  - Multiple Choice Questions On Solubility Curve Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Choice Questions On Solubility Curve
  - Highlighting and Note-Taking Multiple Choice Questions On Solubility Curve
  - Interactive Elements Multiple Choice Questions On Solubility Curve
8. Staying Engaged with Multiple Choice Questions On Solubility Curve
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Multiple Choice Questions On Solubility Curve
9. Balancing eBooks and Physical Books Multiple Choice Questions On Solubility Curve
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Choice Questions On Solubility Curve
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Questions On Solubility Curve
  - Setting Reading Goals Multiple Choice Questions On Solubility Curve
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Questions On Solubility Curve
  - Fact-Checking eBook Content of Multiple Choice Questions On Solubility Curve
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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