

9. Give two example for each of the following
- Strong acid
 - Weak acid
 - Strong base
 - Weak base
 - Conjugate acid/base
 - Buffer
 - Ka
 - pKa
 - pH
 - Kw
 - pKw

10. Calculate $[H^+]$, $[OH^-]$, pH for a 0.015 M aqueous solution of HCl.

11. Phosphoric acid has three pK_a values:

a. Find the form which is most abundant with the ionization equation corresponding each of the pK_a values.

b. buffer capacity is the ability of a solution to resist changes in pH upon addition of strong acid or base. Which of the phosphate ions is the best buffer in its highest concentration at pH=7?

12. Calculate $[H^+]$, $[OH^-]$, pH for a 0.45 M aqueous solution of NaOH.

c. Which of the phosphate ions is the best buffer in its highest concentration at pH=7?

13. Calculate $[H^+]$, $[OH^-]$ and

pH for a 0.15 M aqueous solution of Acetic acid which has pK_a of 4.76.

Abbrevation Of Pkw In Buffer Solution

Christian Drosten



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