

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. acid/base | 6. equilibrium constant |
| | 7. weak solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include correct units when appropriate.

1. Water autoionizes. K_w is the ion-product constant for water.
2. The autoionization of water is $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
3. The autoionization of water is $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
4. The autoionization of water is $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
5. The autoionization of water is $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
6. The autoionization of water is $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
7. The autoionization of water is $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
8. The autoionization of water is $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
9. The autoionization of water is $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.
10. The autoionization of water is $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$.

1. K_w
2. K_w
3. K_w
4. K_w
5. K_w
6. K_w
7. K_w
8. K_w
9. K_w
10. K_w

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