

6.14. Density:

6.14.1. Definition:

1.000 liters of pure water (H_2O) corresponds to 1.000 kg. That is a nice convenience, but does not apply to all other liquids, different in composition from pure water (H_2O): for some liquids, a 1.000 kg of that liquid will weigh more, and for some liquids, a 1.000 kg will weigh less. Obviously, the type of liquid will have an influence, this is how got defined the specific weight, or density of a type of matter/liquid/gas:

$$\text{DENSITY} = \frac{\text{MASS}}{\text{VOLUME}} \quad \text{OR} \quad \rho = m/v$$

Density describes the ratio between the mass of a body of matter, or a liquid on the one hand and the volume it takes up, on the other hand. The symbols and unit of density:

$$\text{DENSITY } [\rho] = \frac{\text{MASS}}{\text{VOLUME}} \quad \frac{[m]}{[V]}$$

$$\text{DENSITY } [\rho] = \frac{\text{MASS}}{\text{VOLUME}} \quad \frac{[kg]}{[m^3]}$$

Density is expressed by the Greek symbol " ρ ", pronounced as "rho". The unit in the metric system is expressed in KILOGRAM/ m^3 and helps to express the mass of 1 m^3 , of a certain substance (solid, liquid, or gaseous). Density is one of the objective criteria to distinguish substances from each other. For example:

EXAMPLE 6.14.1: AvGas (Aviation Gasoline) has a density of 0,72 g/cm³.

How many kg/m³ does that correspond with?

- a. 72.000 kg/m³
- b. 720 kg/m³
- c. 72 kg/m³
- d. 7.200 kg/m³



ANSWER:

Here we are facing a conversion of 2 factors, both the mass and the volume, before we can state anything useful with regards to what was asked.

STEP 1: Let's first tackle the mass, which needs to be converted to the standard unit: g -> kg:

kilogram	hectogram <i>(rarely used unit)</i>	decagram <i>(rarely used unit)</i>	gram	decigram
0,	0	0	7	2

72 g corresponds to 0,0072 kg

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